

ETHIRAJ COLLEGE FOR WOMEN

(AUTONOMOUS)

CHENNAI - 600 008



DEPARTMENT OF CHEMISTRY

Syllabus for

B.Sc. CHEMISTRY

2015 - 2016

**ETHIRAJ COLLEGE FOR WOMEN (AUTONOMOUS)
CHENNAI – 8**

**DEPARTMENT OF CHEMISTRY
BACHELOR OF SCIENCE IN CHEMISTRY
CBC SYSTEM**

(SYLLABUS EFFECTIVE FROM THE ACADEMIC YEAR 2015 -2016)

Department of Chemistry is revising syllabus with effect from the academic year 2015-2016, under CBC system with Part –IV and Part-V as specified by the Government of Tamil Nadu. Part IV and Part V components will seek to build the capacity of the students and provide inputs for their social service and social analysis capabilities.

Every academic year is divided into two semester sessions. Each semester will have a minimum of 90 working days and each day will have 5 working hours.

Teaching is organized into a modular pattern of credit courses. Credit is related to the number of lecture, tutorial and practical hours (LTP) for a particular subject. It is also related to the number of hours a student spends learning a subject or carrying out an activity.

REGULATIONS

1. ELIGIBILITY FOR ADMISSION:

Candidates for admission to the first year of the Degree of Bachelor of Science in Chemistry course shall be required to have passed the Higher Secondary Examinations conducted by the Government of Tamilnadu or an Examination accepted as equivalent thereto by the Syndicate of the University of Madras with Chemistry, Mathematics and Physics as subjects.

2. ELIGIBILITY FOR THE AWARD OF DEGREE:

A candidate shall be eligible for the award of the Degree only if she has undergone the prescribed course of study for a period of not less than three academic years, passed the examinations of all the six semesters prescribed and earned a total of 140 credits including Part-V. Students may earn extra credit by doing Add on Courses.

3. COURSE OF STUDY:

The main subject of study for Bachelor Degree in Chemistry consists of the following.

Study Components	No. of courses	Credit per course	Total credits
PART- I Tamil / Other languages	2+2=4	3	12
PART- II English	2+2=4	3	12
PART- III Core subject :			
Theory	10	-	43
Practicals	5	-	17
Elective	3	5	15
Allied subject:			
Theory	2+2=4	10+8	18
Practicals	1	2	2
PART - IV			
1. (a) Not studied Tamil upto XII Std. -shall take Tamil comprising of two courses (level 6 th Std.)	1+1=2	2	4
(b) Studied Tamil up to XII Std.- taken non-Tamil under Part - I shall take Advanced Tamil comprising of two courses.			
(c) Others who do not come under (a & b) can choose non-Major Elective comprising of two courses.			
2. Soft Skills	4	3	12
3. Environmental studies	1	2	2
4. Value Education	1	2	2
PART - V			
Extension activities	1	1	1
TOTAL			140

4. PASSING MINIMUM:

A candidate shall be declared to have passed in each Theory/ Practical of the main subject of study wherever prescribed, if she secured NOT LESS THAN 40% of the marks prescribed for the end semester examination and also 40% in the sum of end semester examination and continuous assessment. There is no passing minimum for Continuous Assessment (CA).

PART I, II & III:

The maximum marks for CA is 40 and for End Semester examination is 60.
The aggregate mark for each paper is 100.

PART IV:

The maximum marks for the End Semester examination is 50 and the passing minimum is 40% and there is no continuous assessment.

5. CLASSIFICATION OF SUCCESSFUL CANDIDATES:

Part I, II, III & IV

Successful candidates passing the examination and securing the marks

- (i) 60 % and above and
- (ii) 50 % and above but below 60 % in the aggregate shall be declared to have passed the examination in the FIRST CLASS and SECOND CLASS respectively.

All other successful candidates shall be declared to have passed the examination in the THIRD CLASS.

Candidates who pass all the examinations (Part I, II, III and IV) prescribed for the course in the FIRST APPEARANCE ITSELF ALONE are eligible for ranking.

B.Sc. CHEMISTRY
CHOICE BASED CREDIT SYSTEM (CBCS)
From 2015-2016 Batch
COURSE PROFILE

SEMESTER – I						
Course Code	Course Title	Hours/ Week	Credits	Marks		
				CA	SE	Total
Part-I	Tamil/Hindi/French/Sanskrit	5	3	40	60	100
Part-II	English	5	3	40	60	100
Part – III						
CH15/1C/GC1	Core I-General Chemistry I	7	5	40	60	100
	*Core Practical I–Volumetric Analysis	3	-	-	-	-
MA15/1A/AM1	Allied Mathematics-I	6	5	40	60	100
Part-IV						
UG15/1N/BTA UG15/1N/ATA CH15/1N/COS	1a/1b/1c (1a – Basic Tamil , 1b- Advanced Tamil , 1c-Non – Major Elective- Cosmetology)	2	2	-	50	50
	Soft Skill 1	2	3	-	50	50
Total		30	21			

SEMESTER – II

Course Code	Course Title	Hours/Week	Credits	Marks		
				CA	SE	Total
Part-I	Tamil/Hindi/French/Sanskrit	5	3	40	60	100
Part-II	English	5	3	40	60	100
Part –III						
CH15/2C/GC2	Core 2-General Chemistry II	7	5	40	60	100
CH15/2C/PR1	*Core 3 Practical I–Volumetric Analysis	3	4	40	60	100
MA15/2A/AM2	Allied Mathematics-II	6	5	40	60	100
Part-IV						
UG15/2N/BTA UG15/2N/ATA CH15/2N/DPT	1a/1b/1c (1a – Basic Tamil , 1b- Advanced Tamil , 1c-Non – Major Elective – Dyeing and Printing of Textiles)	2	2	-	50	50
UG15/2S/ECS	Soft Skill 2- English Language and Communication Skills II	2	3	-	50	50
Total		30	25			

SEMESTER - III						
Course Code	Course Title	Hours/Week	Credits	Marks		
				CA	SE	Total
Part -I	Tamil/Hindi/Sanskrit/French	5	3	40	60	100
Part-II	English	5	3	40	60	100
Part - III CHI5/3C/GC 3	Core 4-General Chemistry III	7	5	40	60	100
	*Core Practical 2-Inorganic Qualitative Analysis	3	-	-	-	-
PH15/3A/GP1	General Physics-I	4	4	40	60	100
PH15/4A/PPR	*Allied Physics Practicals	2	-	-	-	-
Part-IV	Environmental studies	2	2	-	50	50
	Soft Skill - 3	2	3	-	50	50
Total		30	20			

SEMESTER - IV						
Course Code	Course Title	Hours/Week	Credits	Marks		
				CA	SE	Total
Part-I	Tamil/Hindi/French/Sanskrit	5	3	40	60	100
Part-II	English	5	3	40	60	100
Part – III CH15/4C/GC4	Core 5-General Chemistry IV	7	5	40	60	100
CH15/4C/PR2	*Core 6-Practical 2– Inorganic Qualitative Analysis	3	4	40	60	100
PH15/4A/GP2	General Physics-2	4	4	40	60	100
PH15/4A/PPR	*Allied Physics Practicals	2	2	40	60	100
Part IV	Value Education	2	2	-	50	50
	Soft Skill -4	2	3	-	50	50
Total		30	26			

SEMESTER – V

Course Code	Course Title	Hours/Week	Credits	Marks		
				CA	SE	Total
CH15/5C/INC	Core 7-Inorganic Chemistry	4	4	40	60	100
CH15/5C/ORC	Core 8-Organic Chemistry	4	4	40	60	100
CH15/5C/PHY	Core 9-Physical Chemistry	4	4	40	60	100
CH15/5C/ANC	Core 10-Analytical Chemistry	4	4	40	60	100
CH15/5E/SPE	Elective 1 - Spectroscopy	5	5	40	60	100
	*Core Practical 3 – Organic Chemistry	3	-	-	-	-
	*Core Practical 4 – Gravimetric Analysis	3	-	-	-	-
	*Core Practical 5- Physical Chemistry	3	-	-	-	-
Total		30	21			

SEMESTER – VI						
Course Code	Course Title	Hours/ Week	Credits	Marks		
				CA	SE	Total
CH15/6C/IND	Core 11-Industrial Chemistry	4	3	40	60	100
CH15/6C/PHC	Core 12-Pharmaceutical Chemistry	4	4	40	60	100
CH15/6E/POC	Elective 2 -Polymer Chemistry	5	5	40	60	100
CH15/6E/NAC	Elective 3- Nano Chemistry	5	5	40	60	100
CH15/6C/PR3	*Core 13 – Practical 3-Organic Chemistry	4	3	40	60	100
CH15/6C/PR4	*Core 14 – Practical 4 - Gravimetric Analysis	4	3	40	60	100
CH15/6C/PR5	*Core 15-Practical 5 - Physical Chemistry	4	3	40	60	100
Total		30	26			
Credits at the end of 6 semesters			139			
Part V	NCC/NSS/Sports/CSS	-	1	-	-	-
Total Credits			140			

*Practical Examinations are conducted at the end of Even Semester (II, IV and VI Semesters)

B.Sc. CHEMISTRY
(MAJOR and ALLIED)

(Syllabus effective from the Academic year 2015-2016)

OBJECTIVES OF THE COURSE :

- (i) To provide the basic knowledge in Chemistry and to help students develop analytical skills and attitude.
- (ii) To give the right platform for a future career in Pharmaceutical, Polymer, Nano Chemistry and Industrial Chemistry.
- (iii) Non- Major Electives enable the students to be entrepreneurs.

PREAMBLE:

- Modification of course contents in few courses based on feedback from students and report given by Academic Audit panel members.
- Part IV-Two Non- Major Electives one in each semester for I year , Environmental studies in Semester – III and Value education in semester - IV
- Four Soft Skill Courses in semesters I, II, III & IV
- Practical Examinations are conducted once in an academic year ie., at the end of semester II, IV and VI

COURSE PROFILE:

Total Credits: 140

Semester:	I	II	III	IV	V	VI	NSS/NCC/Sports
Credit :	21	25	20	26	21	26	1

B.Sc Chemistry Major Credits
Total Credits – 75 (Core credits – 60 & Elective Credits – 15)

Sem	Course Code	Part III (CORE) Title of the course	L	T	P	Total Hrs/ Week	Credits	Exam Hrs	Marks		
									CA	SE	Total
I	CH15/1C/GC1	Core 1-General Chemistry -I	4	3	0	7	5	3	40	60	100
II	CH15/2C/GC2	Core 2- General Chemistry -II	4	3	0	7	5	3	40	60	100
	CH15/2C/PR1	Core3- Practical 1-Volumetric Analysis				3	4	3	40	60	100
III	CH15/3C/GC3	Core 4-General Chemistry III	4	3	0	7	5	3	40	60	100
IV	CH15/4C/GC4	Core 5-General Chemistry IV	4	3	0	7	5	3	40	60	100
	CH15/4C/PR2	*Core 6-Practical 2–Inorganic Qualitative Analysis				3	4	3	40	60	100
V	CH15/5C/INC	Core 7-Inorganic Chemistry	3	1	0	4	4	3	40	60	100
	CH15/5C/ORC	Core 8-Organic Chemistry	3	1	0	4	4	3	40	60	100
	CH15/5C/PHY	Core 9-Physical Chemistry	3	1	0	4	4	3	40	60	100
	CH15/5C/ANC	Core 10-Analytical Chemistry	3	1	0	4	4	3	40	60	100
	CH15/5E/SPE	Elective 1 -Spectroscopy	4	1	0	5	5	3	40	60	100
VI	CH15/6C/IND	Core 11-Industrial Chemistry	3	1	0	4	3	3	40	60	100
	CH15/6C/PHC	Core 12-Pharmaceutical Chemistry	3	1	0	4	4	3	40	60	100
	CH15/6E/POC	Elective 2 -Polymer Chemistry	4	1	0	5	5	3	40	60	100
	CH15/6E/NAC	Elective 3- Nano Chemistry	4	1	0	5	5	3	40	60	100
	CH15/6C/PR3	*Core 13 Practical 3 – Organic Chemistry				4	3	3	40	60	100
	CH15/6C/PR4	*Core 14 Practical 4 – Gravimetric Analysis				4	3	3	40	60	100
	CH15/6C/PR5	*Core 15 Practical 5- Physical Chemistry				4	3	3	40	60	100

*Practical Examinations are conducted at the end of Even Semester (II, IV and VI Semesters)

Practical examination for CH15/6C/PR3 & CH15/6C/PR4 will be conducted in the forenoon and afternoon of the same day.

L= Lecture hours, T=Tutorial hours, P=Practical hours

CA= Continuous Assessment marks, SE= End Semester marks

ALLIED CHEMISTRY: Credits: 20 [10 Credits for subjects offered by Department of Chemistry and 10 credits for subjects offered by other departments]

Sem	Course Code	PART III Title of the Course	L	T	P	Total Hrs/ Week	Credits	Exam Hrs	Marks		
									CA	SE	Total
I	CH15/1A/NGC	Allied Chemistry - General Chemistry	3	1	0	4	4	3	40	60	100
II	CH15/2A/NBC	Allied Chemistry -Bio Organic Chemistry	3	1	0	4	4	3	40	60	100
II	CH15/2A/PRA	Allied Chemistry Practical – Volumetric and Organic Analysis				2	2	3	40	60	100
III	CH15/3A/ZGC	Allied Chemistry - General Chemistry	3	1	0	4	4	3	40	60	100
IV	CH15/4A/ZBC	Allied Chemistry- Bio Organic Chemistry	3	1	0	4	4	3	40	60	100
IV	CH15/4A/PRA	Allied Chemistry Practical – Volumetric and Organic Analysis				2	2	3	40	60	100
III	CH15/3A/PGC1	Allied Chemistry –General Chemistry- 1	3	1	0	4	4	3	40	60	100
IV	CH15/4A/PGC2	Allied Chemistry- General Chemistry-2	3	1	0	4	4	3	40	60	100
IV	CH15/4A/PRA	Allied Chemistry Practical – Volumetric and Organic Analysis				2	2	3	40	60	100

NON – MAJOR ELECTIVE: Credits 4

Sem	Course Code	PART IV (Non – Major Elective)	L hr	T hr	P hr	Total Hrs/ Week	Credits	Exam Hrs	End Semester Marks Maximum
I	CH15/1N/ COS	Non – Major Elective -1 - Cosmetology	2	0	0	2	2	2	50
II	CH15/2N/ DPT	Non Major Elective -2 -Dyeing and Printing of textiles	2	0	0	2	2	2	50

The above courses of the UG programme enrich the skills in employability skill development / entrepreneurship which caters the needs of the students.

EVALUATION PATTERN:

Theory: Continuous Assessment (CA)

Pattern for Continuous Assessment

			CA
Test I	2hrs.	50marks	10marks
II	2hrs.	50marks	10
Quiz/ Assignment/ Semester/ Field visit			10
Participatory Learning			10
		Total	40 marks

Practical: Continuous Assessment

(a) One Test	-	10 marks
(b) Skill in practical	-	20 marks
(c) Record	-	10 marks
		Total
		40 marks

End Semester Examination questions are to be chosen from the question bank by the external Examiner and evaluated by both Internal and External Examiners. Duration of Examination is 3 and the maximum mark is 60.

Part IV – Non – Major Elective

- The end semester examination question papers are to be set and evaluated by internal examiner only.
- Duration of examination is 2 hours and the maximum mark is 50.
- No Continuous Assessment.

Rubrics for Continuous Assessment

Assignment: Appearance, contents, originality, presentation, schematic representation and diagram, bibliography.

Seminar: Organization, subject knowledge, visual aids, confidence level, presentation.

Participatory learning: Answering questions, clearing doubts, participation in discussion, attendance, communication and language.

End Semester Examination question papers are to be set by the External Examiners and evaluated by both Internal and External Examiners.

Duration of examination is 3 hours and Maximum mark is 100.

SEMESTER- I

Title of the Course: **Core 1-General Chemistry-I**

Teaching hours: $15 \times 7 = 105$

Course Code: **CH15/1C/GC1**

Credits: 5

L T P 4 3 0

Objectives:

1. To provide basic concepts of Chemical bonding, Inorganic Qualitative and Quantitative Analysis.
2. To provide the basic concepts of Organic Chemistry which will facilitate understanding of reaction mechanisms.
3. To provide the concepts and applications of thermodynamics and a knowledge in thermochemistry, solutions and their behavior.

COURSE OUTLINE

UNIT I: Chemical bonding: ionic bond-Madelung energy-polarization of ions Fajan's rule-covalent bond-Wave mechanical Principles involved in treatment of valence bond and molecular orbital theories-Shapes of molecules. VSEPR BeF₂, BF₃, PF₅, SF₆, IF₇, H₂O, NH₃, IF₅, XeF₂, XeF₄, XeF₆ and XeOF₄, Molecular orbital theory as applied to N₂, F₂, O₂-Super oxo and peroxy ions, CO and NO molecules. Introduction to Band theory- Bonding in metals- **classification as conductors, insulators and semiconductors.**

21 hrs

UNIT II: Principles of qualitative analysis: Laboratory Methods in **Semimicro qualitative Analysis**, Filtration of precipitates, Washing of Precipitates, Heating of Solutions, Evaporation, Methods of Precipitating Sulphides-Types of Reactions Involved in Qualitative Analysis-Dry Reaction-Precipitation Reactions-Complexation Reactions-Preparation of Solution for Cations Testing on Semi Micro Scale. Principles of **quantitative analysis**-Definition-Standardisation-Experimental Requirements of **Volumetric Analysis** Types of Titrations. Indicators for Acid-Base Titrations, Precipitation-Titrations-Redox Titrations-Self Indicators-External Indicators-**Gravimetric analysis** -Conditions for Precipitation-Choice of Precipitants-Types of Organic Precipitants-Specific and selective precipitants-Sequestering Agents-Solubility of Precipitates-Theories of Precipitation-Co-Precipitation-Post Precipitation-Effect of Digestion-Washing of Precipitates-Types and Care of Crucibles.

21hrs

UNIT III: Shapes of molecules: Hybridisation and Geometry of molecule - Methane, Ethane, Ethylene, Acetylene and Benzene. Polar effects- Explanation with examples -

Inductive effect, Inductomeric effect, Electomeric effect, Mesomeric effect, Resonance effect, Hyperconjugation and Steric effects-Steric inhibition of Resonance- Steric retardation and Steric acceleration. Cleavage of covalent bonds-Heterolytic fission-Carbocations, Carbanions-structure and stability. Homolytic fission- free radicals-structure and relative stability of free radicals. **21hrs**

UNIT IV: Thermodynamics I: Thermodynamic functions-state functions and path functions. Exact and inexact differential, Zeroth law of thermodynamics-statement-application, First law of Thermodynamics-statement, equation, Internal energy and enthalpy, relationship between molecular heat capacity at constant volume and constant pressure for an ideal gas-calculation of q , ΔE & ΔH for the expansion of ideal gases under reversible, isothermal and adiabatic conditions. **Thermochemistry**-Enthalpy of reaction, enthalpy of combustion, enthalpy of solutions, enthalpy of dilution, differential & intergral heats of solution, laws of thermo Chemistry, bond energy, calculation of bond energy using thermo chemical data, application of bond energy data, variation of enthalpy of reaction with temperature-Kirchoff's equation. **21hrs**

UNIT V: Solutions: Solutions of gases in liquids, factors influencing solubility of a gas, Henry's law. Solutions of liquids in liquids - Raoult's Law, vapour pressure of ideal solutions, vapour pressure of non-ideal solutions, vapour pressure- composition curves, boiling point-composition curves of completely miscible binary solutions, fractional distillation of completely miscible binary liquid solutions, **azeotropic mixtures**. Solubility of partially miscible liquids pairs, Phenol- water system, Triethylamine water system, nicotine-water system, effect of impurities on CST, distillation of immiscible liquid pairs. **21hrs**

RECOMMENDED TEXTBOOKS:

1. Elements of Analytical Chemistry. R.Gopalan, P.S.Subramanian, K.Rengarajan- 3rd edition
2. Principles of Inorganic Chemistry B.R.Puri L.R. Sharma & K.C.Kalia 28th edn – Vallabh publications.2003
3. Organic Chemistry by P.L.Soni , Sultan Chand and Sons, 28th edition , 2012
4. Organic Chemistry by Bahl & Arun Bahl , 11th edn , Sultan Chand and Co.2010
5. Organic Chemistry by Morrison & Boyd , 7th edn , Pearson India 2011
6. Principles of Physical Chemistry – Puri & Sharma 41st edn Vishal Publishing Co. 2004.
7. Physical Chemistry – Samuel Glasstone, David Lewis, Palgrave Macmillan 1993
8. Guide book to mechanism in Organic Chemistry by Peter Sykes, 6th edition , Pearson Publications, 2003

REFERENCES:

1. Text book of Qualitative chemical analysis 1st edition- G.H.Jeffery, J.Bassett, J.Mendhal, R.C.Deynery.
2. Text book of Quantitative Inorganic analysis-Vogal ELBS III edition, (1976) and IV Edition (1985)
3. Text book of Quantitative Inorganic analysis-Vogal ELBS III edition, (1976) and IV Edition (1985)
4. Organic Chemistry by Finar Vol II 3rd edition ELBS
5. Text book of Physical Chemistry by A.S.Negi & S.C.Anand. 1st edn New Age International.
6. Text book of Physical Chemistry – Kapoor Vol 1,2,3 & 4. Mac Millan
7. Thermodynamics for chemists – Samuel Glasstone, Ist edition East West Press.
9. An Introduction to Chemical Thermodynamics by Rastogi & Misra, 6. th Revised edition Vikas Publishing House.
10. Physical Chemistry by P.W. Atkins, 10 th edition Oxford University Press.

PERIODICALS:

1. Education in Chemistry
2. Indian academy of Sciences-Proceedings-Chemical Sciences
3. Current Science
4. Journal of Inorganic Chemistry
5. Journal of American Chemical Society

WEBSITES & e-LEARNING:

1. www.acs.org
2. www.virtlab.com
3. Internet Chemistry resources
4. nptel.ac.in
5. MATLAB

SEMESTER- II

Title of the Course: Core 2- **General Chemistry-II**

Teaching hours: $15 \times 7 = 105$

Course Code: **CH15/2C/GC2**

Credits: **5**

L T P 4 3 0

Objectives:

1. To provide a knowledge in nuclear Chemistry and nuclear energy
2. To study the reactions of alkanes, alkenes and alkynes and concept of aromaticity
3. To provide the concepts and applications of II and III law of Thermodynamics

UNIT I:

Fundamental particles of nucleus: Nucleon terminology-isobars, isotopes, isomers, mirror nuclei, properties of nucleus - mass, radius, spin - Nuclear forces operative between nucleons. Magnitude of nuclear forces. Stability of nucleus, N/P ratio-curves-stability belts-Nuclear binding energy-mass defect, simple calculations involving mass defect and binding energy per nucleon. Magic numbers, Structure of Nucleus-Liquid drop model-shell model. Artificial transmutation of elements, Induced radioactivity, radio isotopes. **Nuclear fission** by thermal neutrons and **nuclear fusion.** **21hrs**

UNIT II:

Alkanes: Mechanism of free radical substitution of alkanes-halogenations of alkane- reactivity and selectivity. **Alkenes-**Mechanism of Electrophilic and Free radical additions of alkenes-Addition of halogens, hydrogen halides, Markonikoff's rule, Peroxide effect in hydrogen bromide addition- addition of sulphuric acid, water, hydroboration (Addition of diborane to ethylene and propylene only), ozonolysis, hydroxylation with KMnO_4 , allylic substitution by NBS. **Dienes-**Classification, stability and chemical reactivity-1,2 and 1,4- addition to conjugated dienes, kinetic and thermodynamic control of reaction-**Diels-Alder reaction.** **Alkynes-**Acidity of acetylene, Addition of water in presence of HgSO_4 catalyst, halogens, hydrogen halides, **hydroboration** and **ozonolysis.** **21hrs**

UNIT III:

Aromaticity: Huckel's rule of Aromaticity and its simple applications-Benzene, Naphthalene, Anthracene and Phenanthrene. **Polynuclear hydrocarbons-** Naphthalene, Anthracene -synthesis, properties, and uses. **Aromatic electrophilic substitution reaction**-orientation and reactivity- Mechanism of nitration, sulphonation, halogenations, Friedal craft's alkylation and acylation. Nuclear and side chain halogenations. **Aromatic nucleophilic substitution** - Activated aryl halides - Benzyne mechanism -Effect of substituents on reactivity. **21hrs**

UNIT IV: **Thermodynamics II:** Need for second law, statements of second law, concept of entropy-definition-entropy changes in isothermal expansion of an ideal gas, cyclic, reversible and irreversible processes, physical transformation-calculation of entropy changes of an ideal gas with changes in temperature, volume and pressure. Entropy of mixing-Gibbs free energy-Helmholtz free energy- their variation with temperature, pressure and volume. Criteria for spontaneity, Gibbs Helmholtz equation, derivation and applications. Fundamental equations of thermodynamics-Maxwell's relationship- thermodynamic equation of state. **21hrs**

UNIT V: **Partial molar properties:** Partial molar free energy-chemical potential-Gibbs Duhem equation-chemical potential and other thermodynamic functions-relationship with enthalpy, internal energy and work function. Variation of chemical potential with temperature and pressure-chemical potential in a system of pure solid, pure liquid and ideal gas mixture. Gibbs-Duhem Margules equation. Concept of fugacity and activity. Determination of fugacity of gas-activity and activity coefficient. **Thermodynamics III**-Statement-Nernst heat theorem-Evaluation of absolute entropies from heat capacity measurements-exceptions to the third law. **21 hrs**

RECOMMENDED TEXTBOOKS:

1. Organic Chemistry by P.L.Soni , Sultan Chand and Sons, 28th edition , 2012
2. Organic Chemistry by Bahl & Arun Bahl , 11th edn , Sultan Chand and Co. 2010
3. Organic Chemistry by Morrison & Boyd , 7th edn , Pearson India 2011
4. Guide book to mechanism in Organic Chemistry by Peter Sykes, 6th edition , Pearson Publications, 2003
5. Physical Chemistry – Samuel Glasstone, David Lewis, Palgrave Macmillan 1993
6. Principles of Inorganic Chemistry B.R.Puri L.R. Sharma & K.C.Kalia 28th edn – Vallabh publications. 2003
7. Text Book of Physical Chemistry – P.L.Soni – O.P.Dharmarha – U.N.Dash, Sultan Chand and Sons, 22nd revised edition.
8. Advanced Physical Chemistry – J.N.Gurtu and A.Gurtu, A Pragathi edition.
9. Principles of Physical Chemistry – Puri & Sharma 41st edn Vishal Publishing Co. 2000.
10. Atkins, Physical Chemistry, Julio de Paula Peter Atkins, 10th edition, Oxford University press.

REFERENCE BOOKS:

1. Organic Chemistry by Finar Vol II 3rd edition ELBS
2. Text book of Quantitative Inorganic analysis-Vogal ELBS III edition, (1976) and IV Edition (1985)
3. Textbook of Quantitative Inorganic analysis-Vogal ELBS III edition, (1976) and IV Edition (1985)
4. Textbook of Physical Chemistry by A.S.Neg & S.C.Anand. 1st edition, New Age International
5. Text book of Physical Chemistry – Kapoor Vol 1, 2, 3 & 4. Mac Millan
6. Thermodynamics for Chemists–Samuel Glasstone, 2007. D. Van Nostrand Company.

PERIODICALS:

1. Resonance- Journal of science education
2. Education in Chemistry
3. Indian academy of sciences-proceedings- Chemical Sciences
4. Current Science
5. Journal of Indian chemical education
6. Journal of American chemical society

WEBSITES & e-LEARNING SOURCES:

1. www.virtlab.com
2. Internet Chemistry resources
3. www.acs.org
4. nptel.ac.in
5. MATLAB

SEMESTERS I & II

MAJOR PRACTICALS

Title of the Course: **Core 3- Practical 1 - Volumetric Analysis**

Teaching hrs: 30 x 3= 90

Credits: **4**

Course Code: **CH15/2C/PR1**

ACIDIMETRY

1. Estimation of Borax—Standard Sodium carbonate.
2. Estimation of Oxalic acid - Standard Oxalic acid.
3. *Estimation of temporary and permanent hardness of water.

PERMANGANIMETRY

4. Estimation of Ferrous ammonium sulphate-Standard Ferrous sulphate.
5. Estimation of Oxalic acid- Standard Ferrous sulphate.

DICHROMETRY

6. Estimation of Ferrous ion against dichromate using diphenylamine as Internal indicator.

COMPLEXOMETRY

7. Estimation of Magnesium/Zinc using EDTA

***IODIMETRY**

8. Estimation of Arsenious oxide

***IODOMETRY**

9. Estimation of Copper-Standard Potassium dichromate

***CERIMETRY**

10. Estimation of Ferrous ion

***ARGENTIMETRY**

11. Estimation of Chloride by Mohr's method.

***For internal assessment only**

END SEMESTER PRACTICAL EXAMINATION

I B.Sc Chemistry SEMESTERS I & II

QUESTION BANK

Title of the Course: Core 3-Practical 1 -Volumetric Analysis

Time: 3 hrs

Course code: CH15/2C/PR1

Max. Marks: 60

1. You are provided with oxalic acid solution. Estimate volumetrically the amount of oxalic acid present in the whole of the given solution. You are supplied with approximately decinormal solution of sodium hydroxide and pure crystalline oxalic acid. Get your burette and balance readings attested by the examiner.
2. You are given borax solution. Estimate volumetrically the amount of borax present in the whole of the given solution. You are supplied with approximately decinormal solution of hydrochloric acid and pure anhydrous sodium carbonate. Get your burette and balance readings attested by the examiner.
3. You are given ferrous ammonium sulphate solution. Estimate volumetrically the amount of ferrous ammonium sulphate present in the whole of the given solution. You are supplied with approximately decinormal solution of potassium permanganate and pure crystalline oxalic acid. Get your burette readings attested by the examiner.
4. You are given oxalic acid solution. Estimate volumetrically the amount of oxalic acid present in the whole of the given solution. You are supplied with approximately decinormal solution of potassium permanganate and pure ferrous ammonium sulphate crystals. Get your burette readings and the weights attested by the examiner.
5. You are given magnesium sulphate solution. Estimate volumetrically the amount of magnesium sulphate present in the whole of the given solution. You are supplied with approximately decinormal solution of EDTA and pure crystalline zinc sulphate. Get your burette and the weights attested by the examiner.

SCHEME OF VALUATION

Continuous Assessment	:	40 marks
External marks	:	60 marks [50+ 10 marks for Record]
Total	:	100 marks

Volumetric analysis

Error upto 2%	:	50 marks
Error from 2.1% to 3%	:	45 marks
Error from 3.1% to 4%	:	35 marks
Error above 4%	:	20 marks
Arithmetic error	-	reduce 1 mark
Wrong calculation	-	reduce 20% of the marks secured
No calculation	-	reduce 40% of the marks secured

SEMESTER- III

Title of the Course: **Core 4-General Chemistry-III**

Teaching hours: $15 \times 7 = 105$

Credits: **5**

Course Code: **CH15/3C/GC3**

L T P 4 3 0

Objectives:

1. To provide the basic concepts in the field of metallurgy and solid state chemistry.
2. To study the reactions of alcohols, acids and carbonyl compounds.
3. To provide the concepts and applications of Nernst distribution law and colligative properties.

COURSE OUTLINE

UNIT I: Basic metallurgical processes: ores, minerals-Metal ore mines in India, General methods of extraction-concentration methods, reduction methods, refining processes. Extraction of Ti, V, Cr, Ni from their principle ores, preparation and properties of compounds of Ti, V, Cr, Ni, Co-oxides, halides, oxy halides and metal salts.

21hrs

UNIT II: Solid state: Introduction-Elements of Symmetry-Symmetry operations. Classification into groups-Point groups of molecules like H_2O , NH_3 only. (Restrict to C_{2v} , C_{3v} & C_{2h})-Space lattice, Unit lattice, packing- hcp & ccp, Bravais Lattices-Different crystal systems, Miller indices, Designation of planes, X-ray diffraction, **Bragg's equation**, rotating crystallography method, **powder diffraction**-structure determination. Structure of rock salt. Spalerite, Wurtzite, Cesium Chloride Structures. Schottky and Frenkel defects.

21hrs

UNIT III: Phenols and Alcohols: Acidic character of Phenols & Alcohols-Explanation on the basis of resonance stabilization. Ring substitution in phenols-orientation of phenolic group towards electrophiles. Mechanism of esterification, nitration, sulphonation, halogenation, coupling with diazonium salts. Kolbe's reaction, Reimer Tieman reaction, Gattermann reaction, Houben Hoesch reaction and Phthalein fusion reaction. Alpha and beta naphthols-preparation, properties and uses. **Unsaturated alcohols**-Preparation and reactions of allyl alcohol. Aliphatic nucleophilic substitution reaction-Mechanism and stereochemistry of SN_1 , SN_2 , SN_i reaction- effect of solvents, leaving groups, nucleophiles and substrates.Elimination reaction-E1 and E2 eliminations-Mechanism and stereochemistry, dehydration of alcohols and dehydrohalogenation-**Hoffmann and Saytzeff's rule.**

21hrs

UNIT IV: Carbonyl Compounds: Polarisation and acidity of alpha hydrogen in carbonyl compounds. Mechanism of Nucleophilic addition-Reactions of Carbonyl compounds with NaHSO₃, HCN, RMgX, H₂N-NH₂, RNH₂, C₆H₅NH₂, phenyl hydrazine and 2,4-dinitro phenyl hydrazine. Mechanisms of Aldol and Benzoin condensations, Perkin, Knoevenagel, Claisen, Wittig, Cannizzaro and Reformatsky reactions. Mechanism of Reduction of carbonyl compounds with sodium borohydride, LiAlH₄, Wolf-Kishner, Clemmenson reduction and MPV reduction. **Grignard reagent-** Preparation and synthetic applications only. Dicarboxylic Acids- Preparation and properties of oxalic, Malonic, Succinic, Glutaric and Adipic acids. Unsaturated dicarboxylic acid-Preparation and properties of Maleic and Fumaric acid. Characteristics reactions of Active methylene group- preparation and synthetic uses of Malonic ester, Acetoacetic esters and cyanoacetic esters. **Tautomerism-**definition- Keto-enol tautomerism (Identification, acid and base catalysed inter conversions, mechanism, preparations and characteristics). **21hrs**

UNIT V: Nernst distribution law: validity of the law, thermodynamic derivation, deviation from distribution Law-association, dissociation and chemical combination of the solute with one of the solvents. **Applications-**association and dissociation of a solute, complex ions and solvent extraction. **Colligative properties-** Osmosis and Osmotic pressure, Determination of molar mass of the solute from vapour pressure lowering, Elevation in boiling point, Depression in freezing point and Osmotic pressure measurements. Vant Hoff factor and deviations. **21hrs**

RECOMMENDED TEXT BOOKS:

1. Principles of Inorganic Chemistry B.R.Puri L.R. Sharma & K.C.Kalia 28th edn – Vallabh publications, 2003
2. Organic Chemistry by P.L.Soni , Sultan Chand and Sons, 28th edition , 2012
3. Organic Chemistry by Bahl & Arun Bahl , 11th edn , Sultan Chand and Co.2010
4. Organic Chemistry by Morrison & Boyd , 7th edn , Pearson India 2011 .
5. Text book of Physical Chemistry by Puri & Sharma 41st edn Vishal Publishing Co. 2003.
6. Physical chemistry by Samuel Glasstone, David Lewis, Palgrave Macmillan 1993
7. Text Book of Physical Chemistry – P.L.Soni – O.P.Dharmarha – U.N.Dash, Sultan Chand and Sons, 22nd revised edition.
8. Guide book to mechanism in Organic Chemistry by Peter Sykes, 6th edition , Pearson Publications, 2003
9. Introduction to solids, L.V. Azaroff MacGraw Hill New york 1960
10. Solid State Chemistry and its Applications, 2nd Edition, Student Edition, Anthony R. West, 2014 Wiley

REFERENCE BOOKS:

1. Text book of Quantitative Inorganic analysis-Vogal ELBS III edition, (1976) and IV Edition (1985)
2. Organic Chemistry by Finar Vol II 3rd edition ELBS
3. UG-Organic Chemistry, Volume II, Jagadamba Singh & L.D.S. Yadav, 8th edition 2013.

PERIODICALS:

1. Resonance- Journal of science education
2. Education in Chemistry
3. Indian academy of sciences-proceedings- Chemical Sciences
4. Current Science
5. Journal of Indian chemical education
6. Journal of Inorganic Chemistry.
7. Journal of American chemical society

WEBSITES & e-LEARNING SOURCES:

1. www.acs.org
2. www.virtlab.com
3. Internet Chemistry resources
4. nptel.ac.in
5. MATLAB

SEMESTER- IV

Title of the Course: **Core 5-General Chemistry-IV**

Teaching hours: $15 \times 7 = 105$

Credits: **5**

Course Code: **CH15/4C/GC4**

L T P 4 3 0

Objectives:

1. To study the reactions of Amines, Nitro compounds and chemistry of natural products
2. To understand the structures of common inorganic polymers, their preparation and applications.
3. To provide the concepts and applications of Kinetics, Photo Chemistry and Surface Chemistry

COURSE OUTLINE

UNIT I: Amines: Basicity of amines, Reactions of amines- Carbylamine reaction. **Nitro Compounds-** Conversions of nitrobenzene to ortho, para, meta dinitrobenzene, TNT. Aromatic nitrocompounds-Reduction in neutral, acidic and alkaline media. Diazotisation and its mechanism. Synthetic application of diazonium salts. Diazomethane and diazoacetic esters-preparation, structures and their synthetic uses. **Dyes-**Theory of colour and constitution, classification-according to structure and method of application. **Preparation and uses** of 1.Azo dye-Methyl orange, 2.Triphenyl methane dye-Malachite green, 3.Phthalein dye-Phenolphthalein and fluorescein, 4. Vat dye-Indigo, 5.Anthraquinone dye-Alizarin **21 hrs**

UNIT II: Carbohydrates: **Carbohydrates-classification**, constitution of glucose and fructose. Reactions of glucose and fructose-osazone formation, mutarotation and its mechanism-cyclic structures-pyranose and furanose forms. Haworth projection formula, Epimerisation. Configuration of monosaccharides -chain lengthening and chain shortening of aldoses, **Interconversion** of aldoses and ketoses. Disaccharides- Properties and structure of sucrose & maltose (Structural elucidation not necessary). Polysaccharides- Properties and structure of starch and cellulose (Structural elucidation not necessary). **Aminoacids and Proteins:** Aminoacids-Classification of amino acids-essential and non essential aminoacids-Preparation of alpha amino acids-from halogenated acids, Gabriels phthalimide synthesis, Strecker synthesis. Properties and reactions-**Zwitter ions**, isoelectric points - peptide synthesis (Bergmann's method only)-structure determination of polypeptides-endgroup analysis. Proteins-Classification based on physical and chemical properties and physiological function-**Primary and secondary structures of proteins**-Helical and sheet structures (Elementary treatment only)-Denaturation of proteins-Colour reaction of proteins. Nucleic acids-Nucleoside, Nucleotides- functions of Nucleotides, Nucleotides as energy carriers, types of Nucleic acid, Structure and functions of nucleic acids. **21hrs**

UNIT III: Synthesis and Applications of Polymers Containing Inorganic Atoms: Main group containing polymers: polysiloxanes, polysilanes, polyphosphazenes. Polysilicates-classification, preparation and applications. Transition metal containing polymers. Sulphur nitrogen polymers, Boron based polymers. General properties of Inorganic polymers-Glass transition temperature- phosphorous based polymers-Maddrell's salts-coordination polymers with two and three dimensional network-compounds-Preparation properties and uses. **Ziegler-Natta catalysis:** used to prepare polyethylene and polypropylene. **21hrs**

UNIT IV: Chemical Kinetics: Rate of a chemical reaction-factors affecting the rate of a chemical reaction- order and molecularity of a reaction. Derivation of rate constants for zero, first, second and third order reactions (equal concentration of reactants) derivation of time for half change- methods of determining order of a reaction- half life method, experimental methods involved in the study of kinetics- Volumetry, Polarimetry and Colorimetry. Chemical Kinetics II-Complex reactions-reversible, parallel and consecutive reactions (only examples). Effect of temperature on the rate of reactions-Arrhenius equation, concept of energy of activation- calculation of Arrhenius parameters. Collision theory and derivation of rate constant for bimolecular reaction- Lindemann's theory of unimolecular reactions- **Absolute Reaction rate theory (ARRT).** Thermodynamic derivation of rate constant for a bimolecular reaction. Comparison of collision theory with ARRT- Significance of entropy and free energy of activation. **21hrs**

UNIT V: Photochemistry: Consequences of light absorption, Jablonski diagram-radiative and non radiative transitions, laws of photochemistry, quantum yield and determination-actinometry. Kinetics of hydrogen-chlorine reaction, kinetics of hydrogen-bromine reaction. **Photosensitization, Chemiluminescence.** **Surface Chemistry-** Physisorption, Chemisorption, Langmuir isotherm, Freundlich isotherm, BET isotherm (no derivation) surface area determination. **Applications of adsorption** (general aspects only). Catalysis-types, classification, characteristics. Theories of catalysis-Homogenous catalysis-Acid-Base, Intermediate and compound formation. Gibbs adsorption isotherm-derivation, application. Surface films. **21 hrs**

RECOMMENDED TEXTBOOKS:

1. Principles of Inorganic Chemistry B.R.Puri L.R. Sharma & K.C.Kalia 28th edn- Vallabh, publications. 2003 Recommended Books.
2. Polymers, D. Walton and P. Lorimer, 2001, Oxford Chemistry Primers
3. Inorganic Polymers, J. E. Mark, H. R. Allcock and R. West, 1992, Prentice Hall.
4. Principles of Physical Chemistry Puri and Sharma and Pathania 2003 Millenium edition, Vishal Publishing Co.
5. Undergraduate Physical Chemistry, Gurtu. Gurthu Volume I, II, III 5th edition (2013), Pragathi Prakashan
6. Text Book of Physical Chemistry – P.L.Soni – O.P.Dharmarha – U.N.Dash, Sultan Chand and Sons, 22nd revised edition.
7. Fundamentals of Biochemistry for Medical students – Ambika Shanmugam, 2012, Lippincott Williams & Wilkins publications.
8. Organic Chemistry by Jagdamba Singh, L.D.S. Yadav 12th edition 2014
11. Organic Chemistry by P.L.Soni , Sultan Chand and Sons, 28th edition , 2012
12. Organic Chemistry by Bahl & Arun Bahl , 11th edn , Sultan Chand and Co. 2010
13. Organic Chemistry by Morrison & Boyd , 7th edn , Pearson India 2011
14. Guide book to mechanism in Organic Chemistry by Peter Sykes, 6th edition , Pearson Publications, 2003

REFERENCES:

1. Introductory polymer chemistry – G.S.Mishra New Age 2007
2. Polymer science – V.R.Gowriker, N.V.Viswanathan and Jayadev sreedar New Age 2006
3. Principles of polymer science –P.Bhadur, N.V.Sastry, Alpha Science International Ltd 2005
4. Physical Chemisry, Peter Atkins. Julio De Paula, Ninth edition (2010), Oxford University press.
5. Organic Chemistry by Finar Vol II 3rd edition ELBS
6. Text book of natural products by chatwal –vol I & vol II Himalayan Publications
7. Text book of Biochemistry - Agarwal, Goel Publications.
8. Fundamentals of Biochemistry – J. L. Jain, Sultan Chand and Sons.

PERIODICALS:

1. Resonance- Journal of science education
2. Education in Chemistry
3. Indian academy of sciences-proceedings- Chemical Sciences
4. Current Science
5. Journal of Indian chemical education
6. Journal of Polymer Science
7. Journal of Polymers
8. Journal of Polymer Research

WEBSITES & e-LEARNING SOURCES:

1. www.acs.org.
2. www.virtlab.com
3. Internet chemistry resources.

SEMESTERS – III & IV
MAJOR PRACTICALS

Title of the Course: **Core 6- Practical 2- Inorganic Qualitative Analysis**

Course Code: **CH 15/4C/PR2**

Credits: **4**

Teaching hours: **30 x 3 = 90**

Analysis of a mixture containing two anions one of which is interfering and two cations.

Anions:

Non interfering Cl^- , Br^- , CO_3^{2-} , SO_4^{2-} and NO_3^-

Interfering PO_4^{3-} , BO_3^{3-} , AsO_3^{3-} , F^- , $\text{C}_2\text{O}_4^{2-}$ and CrO_4^{2-}

Cations:

Pb^{2+} , Cu^{2+} , Bi^{3+} , Cd^{2+} , Fe^{2+} , Fe^{3+} , Al^{3+} , Co^{2+} , Ni^{2+} , Zn^{2+} , Mn^{2+} , Ba^{2+} , Ca^{2+} ,
 Sr^{2+} , Mg^{2+} and NH_4^+

End Semester Practical Examination
II B.Sc Chemistry-Semesters-III & IV

QUESTION BANK

Title of the Course: Core 6-Practical 2-Inorganic Qualitative Analysis

Course Code: CH15/4C/PR2

Time: 3hrs
Max. Marks: 60

Analyze systematically the given mixture and report for one interfering anion, one non- interfering anion and two cations

SCHEME OF VALUATION

Total Marks = 100 (CA 40+End Semester 60)

Each metal ion	24 marks	(2 x 12 marks)
Non- Interfering anion	12 marks	
Interfering anion	14 marks	
Record	10 marks	
Total	60	

SEMESTER – V

Title of the course: **Core 7- Inorganic Chemistry**

Teaching hours: 15 x 4 = 60

Credits: **4**

Course Code: **CH15/5C/INC**

L T P 3 1 0

Objectives:

1. To focus on the basics of co-ordination chemistry, organometallic Chemistry, Solid state and Data Analysis.
2. To concentrate on Bio-Inorganic chemistry stressing on the role of metal ions in biological system.
3. To understand the method, isolation & separation of lanthanide elements.

COURSE OUTLINE

UNIT I: Introduction to coordination compounds: ligand classification, nomenclature of coordination compounds, isomerism in complexes-ionization, hydrate, position, linkage, geometrical, and optical isomerisms. Thermodynamic stability of complexes. Determination of stability constant by Job's method. Factors affecting stability of a complex– **chelate effect.** **12hrs**

UNIT II: Coordination Chemistry: **Theories of coordination-Werner, Sidgwick, EAN rule, Pauling's valence bond theory,** Electroneutrality principle, Pi bonding concept, Inner & outer orbital complexes-Merits & Demerits. Crystal Field Theory, Octahedral and tetrahedral splitting, 10 Dq, Stabilization energy- CFSE. Measurement of 10 Dq- Factors affecting 10 Dq-Spectrochemical Series-**Jahn Teller effect.** **12hrs**

UNIT III: Oxygen carriers: **Haemoglobin, Myoglobin,** Bohr effect, Perutz mechanism-oxygen dissociation curves-**Vitamin B12**-Role of Cobalt-Zinc containing enzymes-Carboxy peptidase, Carbonic anhydrase. Function of Na⁺/K⁺ Mechanism of ion pump-Role of Calcium. **Role of Copper in ascorbic acid oxidase, Function of ceruloplasmin,** Transamination reactions- Role of Magnesium in Hexokinase. **12hrs**

UNIT IV: Inner transition metals: lanthanides- Extraction, Separation techniques, Extraction of thorium and uranium- **Artificial synthesis of trans uranium elements-** Significance of rare earths-Electronic Configurations-Common oxidation states-Lanthanide Contraction-consequences. **12hrs**

UNIT V: Organo metallic chemistry: Ferrocene-Preparation Reactivity. Carbonyls and nitrosyls Structure and bonding-Preparation Reactivity. Alkene and alkyne complexes. π bonding. Zeiss salt. Organo lithium and Boron complexes. **12hrs**

RECOMMENDED TEXTBOOKS:

1. Principles of Inorganic chemistry B.R.Puri, L.R.Sharma & K.C.Kalia 28th edn, Vallabh Publications, 2003.
2. Co-ordination Chemistry, D. Banerjea, Tata McGraw Hill, 1993.
3. Vogel's text book of quantitative Inorganic Analysis – Longman press.
4. Inorganic Chemistry, J.E.Huheey, Harper & Collins NY, IV edn.

REFERENCE BOOKS:

1. Inorganic Chemistry D.F.Shriver, P.W.Atkins, Oxford University Press, 3rd Edn. 1999.
2. Inorganic Chemistry K.F.Purcell & J.C.Kotz Swb Saunders Co. 1977.
3. Co-ordination Chemistry - SFA Kettle, ELBS, 1973.

PERIODICALS:

1. Resonance- Journal of science education
2. Education in Chemistry
3. Indian academy of sciences-proceedings - Chemical Sciences
4. Current Science
5. Journal of Indian chemical education
6. Journal of American chemical society

WEBSITES & e-LEARNING SOURCES:

1. www.acs.org
2. www.virtlab.com
3. Internet chemistry resources

SEMESTER- V

Title of the Course: **Core 8-Organic Chemistry**

Teaching hours: $15 \times 4 = 60$

Credits: **4**

Course Code: **CH15/5C/ORC**

L T P 3 1 0

Objectives:

1. To understand the concepts of stereochemistry and conformational analysis of open chain and cyclohexane ring systems
2. To study the synthesis and reactions of heterocyclic compounds, and molecular rearrangements.
3. To study the chemistry of natural products- alkaloids, terpenes

COURSE OUTLINE

UNIT I: Stereoisomerism: Definition-classification into optical and geometrical isomerisms. Geometric isomerism-cis and trans isomers, syn and anti and E, Z-nomenclature, Methods of assigning configuration, Auwer's skita rule. Optical isomerism-Optical activity-Specific rotations. Conditions for optical activity-assymentrical center-chirality-chiral and achiral molecules-meaning of + and-/ d and l notations- Elements of symmetry-Racemisation- methods of racemisation and mechanism of racemisation-Resolution- methods of resolution(mechanical, seeding, biochemical and conversion into diastereoisomers)- Assymmetric synthesis-partial and asymmetric synthesis, Walden inversion, Projection formulae- Fischer, Flying Wedge, Saw horse and Newmann Projection formulae- Configurational notations of optical isomers-D and L notations, Cahn Ingold and Prelog rules-R,S-notations of optical isomers with one and two asymmetric carbons-Erythro and Threo representations. **12hrs**

UNIT II: Heterocyclic Compounds: Aromaticity of heterocyclic compounds - Synthesis and reactions of Pyrrole, Furan, Thiophene, Pyridine, Quinoline, Isoquinoline and Indole with special reference to Skraup, Bischler, Napieralski and Fischer-indole synthesis. **12hrs**

UNIT III: Conformationl Analysis: Introduction of the terms conformers, configuration, dihedral angle, torsional strain and conformational analysis. Conformational analysis of ethane and n-butane with energy diagrams Conformers of Cyclohexane-boat, chair and skew boat forms-axial and equatorial bonds-ring flipping. **12hrs**

UNIT IV: Molecular Rearrangements: Classification as anionotropic, Cationotropic-intermolecular and intramolecular rearrangement. Mechanism of rearrangements- Pinacol-pinacolone, Beckmann, Benzidine, Hoffmann, Curtius, Benzil-Benzilic acid and Claisen rearrangements. **12hrs**

UNIT V: Natural Products: Terpenoids-classification, isoprene rule, Camphor and Alpha Pinene-structural elucidation. Alkaloids-general methods of isolation and general methods of structure determination of Conine, Piperine and Nicotine. Vitamins- Thiamine, Riboflavin, Pyridoxine and Ascorbic acid-Structure, occurrence and biological importance- structural elucidation of Ascorbic acid only. **12hrs**

RECOMMENDED TEXTBOOKS:

1. Stereochemistry: Conformation and Mechanism by P.S.Kalsi , 7th New Age International Pvt Ltd Publishers , 2015
2. Organic Chemistry by Morrison & Boyd ,7th edn , Pearson India 2011
3. Organic chemistry by Finar Vol I ,Pearson Education; 5 edition (2002)
4. Organic chemistry by Finar Vol II, Pearson Education; 5 edition (2002)

REFERENCES BOOKS:

1. Stereochemistry of carbon compounds by L.Elviel Mac Graw Hill
2. Chemistry of Natural Production by Gurdeep .R.Chatwal. Himalaya Publishing
3. UG-Organic Chemistry, Volume I, Jagadamba Singh & L.D.S. Yadav, 8th edition 2013.

PERIODICALS:

1. Resonance- Journal of science education
2. Education in Chemistry
3. Indian academy of sciences-proceedings- Chemical Sciences
4. Current Science
5. Journal of Indian chemical education
6. Journal of American chemical society

WEBSITES & e-LEARNING SOURCES:

1. www.acs.org
2. www.virtlab.com
3. Internet chemistry resources
4. *nptel.ac.in*
5. *MATLAB*

SEMESTER - V

Title of the Course: **Core 9- Physical Chemistry**

Teaching hours: 15 x 4 = 60

Credits: **4**

Course Code: **CH15/5C/PHY**

L T P 3 1 0

Objectives:

1. To provide knowledge in electrochemistry and its applications
2. To understand phase equilibria and its applications
3. To learn the basics of quantum chemistry

COURSE OUTLINE:

UNIT I: Electrochemistry I: Conductance in electrolytic solutions, Cell constant-Measurement cell constant, equivalent conductivity, Kohlrausch's law and its application-relation between specific, equivalent and molecular conductivity, Effect of dilution on Conductivities. Ionic mobility, transport number, determination by Hittorf's and Moving boundary method, **Debye-Huckel Onsager theory**, equation (no derivation), verification. Debye Falkenhagen, Debye-Wein effect, Debye-Huckel limiting law (no derivation). Applications of conductivity measurements-determination of pH, K_a , K_{sp} , **conductometric titrations**-acid-base titrations and precipitation titrations. **12 hrs**

UNIT II: Electrochemistry II: **Galvanic cells**- reversible and irreversible cells, e.m.f and its measurements- standard cell. **Types of reversible electrodes**-metal/metal ion, metal amalgam/metal ion, metal insoluble salt/anion, gas electrode, redox electrode. Electrode reaction, measurement of electrode potentials using reference electrodes-Standard hydrogen electrode, calomel electrode. **Derivation of Nernst equation for emf** of cells and electrode potentials-sign convention, sign of e.m.f and spontaneity of a reaction, calculation of ΔG and ΔH from emf data. Chemical cells with and without transport, concentration cells with and without transference. **12 hrs**

UNIT III: Electrochemistry III: Application of Gibbs Helmholtz equation in the calculation of thermodynamic quantities in Galvanic cells. Determination of pH using quinhydrone electrode and glass electrode. **Potentiometric titrations**- acid.-base titrations, redox titrations, precipitation titrations, valency of ions, transport-number, K_{sp} and activity coefficient of electrolytes. Irreversible electrode processes- polarisation and over voltage, decomposition voltage. Determination of decomposition voltage and over voltage. **Corrosion**-mechanism, types and methods of prevention. **12hrs**

UNIT IV: Phase Equilibria: Definition of terms-phase, component, degrees of freedom, Gibbs phase rule (Statement only). Applications to one component systems-water, sulphur systems. Two component system- solid-liquid equilibria. Simple eutectic-Lead-silver system. Compound formation with congruent melting point-Magnesium-zinc system. Incongruent melting point-sodium-potassium system. Salt-hydrates- KI- water, FeCl₃-water, freezing mixtures. **12 hrs**

UNIT V: Quantum Chemistry: Failures of classical theory, Black Body radiation, quantum theory of radiation, photoelectric effect, dual character of electron-de Broglie equation-Davison and Germer experiment. Bohr's model atom (no derivation), spectrum of hydrogen atom, Compton effect, Heisenberg uncertainty principle, postulates of quantum mechanics, Schrodinger wave equation (no derivation), significance of Ψ and Ψ^2 Eigen values and eigen functions. **12 hrs**

RECOMMENDED TEXT BOOKS:

1. Text book of Physical Chemistry by Puri & Sharma 41st edn Vishal Publishing Co. 2004
2. Physical chemistry by Samuel Glasstone , David Lewis, Palgrave Macmillan 1993
3. Text book of Physical Chemistry by A.S.Negi & S.C.Anand. 1st edn New Age International
4. Text book of Physical Chemistry - Kapoor Vol 1,2,3 & 4. MacMillan
5. Text Book of Physical Chemistry – P.L.Soni-O.P.Dharmarha – U.N.Dash, Sultan Chand and Sons, 22nd revised edition.
6. Advanced Physical Chemistry – J.N.Gurtu and A.Gurtu, A Pragathi edition.

REFERENCE BOOKS:

1. Electrochemistry – Samuel Glasstone
2. Phase rule and Phase Reactions by Sydney Bowden. Mac Millan 1950

PERIODICALS:

1. Resonance- Journal of science education
2. Education in Chemistry
3. Indian academy of sciences-proceedings- Chemical Sciences
4. Current Science
5. Journal of Indian chemical education
6. Journal of American chemical society

WEBSITES & e-LEARNING SOURCES:

1. www.acs.org
2. www.virtlab.com
3. Internet chemistry resources

SEMESTER-V

Title of the Course: **Core 10-Analytical Chemistry**

Teaching hours: 15 x 4 = 60

Credits: **4**

Course Code: **CH15/5C/ANC**

L T P 3 1 0

Objectives:

1. To impart knowledge and skill in data analysis, separation techniques,
2. To expose to electro analytical techniques
3. To introduce computer applications in chemistry.

COURSE OUTLINE

UNIT I : **Data Analysis:** Theory of **errors**-idea of significant figures and its importance with examples. **Precision accuracy**-methods of expressing accuracy-error analysis, minimizing errors, methods of expressing precision-**average deviation-standard deviation** and confidence limit. Reporting of data, Presentation of tabulated data, scatter diagram, Method of **least squares**, correlation for straight lines only. **12hrs**

UNIT II: **Thermal analytical methods:** Principle - instrumentation and application of **TGA, DTA** in **Thermogravimetric** and Differential thermal analysis techniques (AgNO_3 , $\text{CaC}_2\text{O}_4 \cdot \text{H}_2\text{O}$, $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ only)-**Thermometric titration**. DTA in different atmospheres - DTG curves ($\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ only). **12hrs**

UNIT III: **Electro analytical methods:** **Polarography** - concentration polarization-DME-convection, migration and diffusion current-**Ilkovic equation** (derivation not required). Experimental assembly-electrodes-solutions current voltage curve - Effect of oxygen -**Applications**. **12hrs**

UNIT IV: **Separation and Purification Techniques:** **Separation techniques**-principles involved in the separation of precipitates-**solvent extraction**. Purification of solid organic compounds-solvent extraction. **Chromatography**-Principles of adsorption, column chromatography-TLC and **paper chromatography**- R_f values-applications. **Ion exchange Chromatography**-principle experimental Techniques-applications-separation of Zn-Mg, Co-Ni, Cd-Zn and chloride-Bromide. **12hrs**

UNIT V: **Introduction to computers & its Applications:** Introduction to Computers-Characteristics of a Computer, **Types of Computers**, Block Diagram of a digital Computer Input & Output devices, memory Storage Systems. **Central Processing Unit**, Binary Numbers, Floating Points, Computational errors. Computer viruses **high level and low level languages**. Basic Introduction to C, Importance of C,

Sample C programs, Basic Structure of C programs. Programming Style- Executive C program., algorithmic flow chart, Character set, C tokens Keywords and Identifiers Variable constant and data types. Declaration of variables. Deferring Symbolic Constants (Basic Constants Only). **Application of C program in Chemistry**: Determination of normality, Molarity and Molality of solutions, Calculation of pH, Internet Service- Provides in India, Terms used in internet www.http, html, URL, TCP/IP bandwidth, dial-up services, ISDN. E-mail-free e-mail providers, search engines uses of internet for chemistry, data base. **12hrs**

RECOMMENDED TEXT BOOKS:

1. Elements of Analytical Chemistry –R.Gopalan,P.S.Subramanian and K.Rengarajan Sultan Chand and Sons.
2. Instrumental methods of analysis by H.Willard, W.Meritt, J Dean. 6th edn Van Nostrand 1981
2. Computers and their applications to chemistry Ramesh Kumari, Narosa.
3. Text book of quantitative inorganic analysis- Vogel ELBS III Edition, (1976) and IV Edition (1985)
4. Polarography by Kapoor Sultan and Chand
5. ANSCIC by Balgurusamy

PERIODICALS:

1. Resonance- Journal of science education
2. Education in Chemistry
3. Indian academy of sciences-proceedings- Chemical Sciences
4. Current Science
5. Journal of Indian chemical education
6. Journal of American chemical society

WEBSITES & e-LEARNING SOURCES:

1. www.acs.org
2. www.virtlab.com
3. Internet chemistry resources

SEMESTER-V

Title of the Course: **Elective 1- Spectroscopy**

Teaching hours: 15 X 5 = 75

Credits: **5**

Course Code: **CH15/5E/SPE**

L T P 4 1 0

Objective:

1. To introduce spectroscopy at basic level
2. To understand the principle and instrumentation
3. To study the applications of spectroscopic techniques

COURSE OUTLINE

UNIT I: Infrared Spectroscopy: origin of **IR spectroscopy**-Simple Harmonic oscillator model of vibrator-Zero point energy-Force constant & bond strengths-Anharmonicity-Morse potential energy diagram- Energy of vibrational levels. Rotational fine structure, Born-Oppenheimer approximation, normal modes of vibration, fundamentals, overtones, hot bands, selection rules. **Instrumentation**, single and double beam instrument- **Applications**-Organic structural elucidation-simple problems, Hydrogen bonding, kinetic studies, identifying cis-trans isomers, linkage isomers, tautomers, impurity detection-quality control-Base line correction method-Quantitative application. **15hrs**

UNIT II: Raman Spectroscopy: Origin-**Stokes and antistokes lines**-Quantum theory- Classical theory. Vibrational Raman spectra-selection rule- polarisability ellipsoids-Changes in polarisability ellipsoids of water and CO₂ during vibrations-plot of polarisability vs displacement coordinate-rule of mutual exclusion **principle-instrumentation**-depolarization ratio-usefulness-connection to Raman spectroscopy-application. **15hrs**

UNIT III: UV-Vis Spectroscopy: Energy levels – molecular orbitals. **Electronic transitions** types- Broad nature of bands. vibrational and rotational transitions accompanying. Selection rule-Effect of Substitution and conjugation-Bathochromic, Hypsochromic, hyper and hypochromic shifts. Instrumentation-Woodward Feiser rules. **Applications**-ualitative and Quantitative analysis. **15hrs**

UNIT IV: NMR Spectroscopy: **¹H NMR only**-Origin-Nuclear Zeeman effect. Instrumentation- Chemical shift.Factors affecting chemical shift. **Spin spin coupling**-fine structure, Pascals triangle-intensity ratio- Fermi contact interaction-Proton count Integration ratio. AB, AX spectrum-Exchange reaction-ethanol spectrum-AMX-**Simple problems**. **15hrs**

UNIT V: Mass Spectroscopy-Basic principles. Theory-Instrumentation-ion production. Molecular ion, Fragment ions, Meta stable peaks-Fragmentation pattern in alkanes, alkenes, alkynes, alcohols, & aromatic alcohols, phenols, Toluene McLafferty rearrangement-nitrogen rule-ring rule. **15hrs**

RECOMMENDED TEXT BOOKS:

1. Applications of absorption spectroscopy of organic compounds by J.Dyer
2. Spectroscopy of Organic chemistry by P.S.kalsi.

REFERENCE TEXTBOOKS:

1. Organic spectroscopy by William Kemp 3rd edn W.H.Freeman & Co,1991
2. Spectroscopic methods in organic chemistry by Silverstein Bassler
3. Instrumental method of analysis by H.Willard, W.Merrit, J Dean. 6th edn Van Nostrand 1981
4. Fundamental of molecular spectroscopy, C.N.Banwell McGraw Hill Newyork 1966

PERIODICALS:

1. Resonance- Journal of science education
2. Education in Chemistry
3. Indian academy of sciences-proceedings- Chemical Sciences
4. Current Science
5. Journal of Indian chemical education
6. Journal of American chemical society

WEBSITES & e-LEARNING SOURCES:

1. www.acs.org
2. www.virtlab.com
3. Internet chemistry resources

SEMESTER- VI

Title of the Course: **Core 11-Industrial Chemistry**

Teaching hours: 15 x 4 = 60

Credits: **3**

Course Code: **CH15/6C/IND**

L T P 3 1 0

Objectives:

1. To expose the students to chemistry in day to day life.
2. To have knowledge in dairy, sugar, paper, leather.
3. To inculcate basic knowledge and awareness in agricultural chemistry

COURSE OUTLINE

UNIT I: Dairy Chemistry: Milk definition, general composition-physico changes taking place in milk due to boiling, pasteurization, sterilization and homogenisation explanation. Components of **milk** – lipids, proteins carbohydrates vitamins, ash and mineral matters names and functions. Definition and compositions of **cream**, butter, ghee, ice-cream, stabilizer and emulsifier. Milk **powder**, definition and need for making manufacture of whole milk powder by **spray drying process**. **12hrs**

UNIT II: Sugar And Paper Industry: Sugar industry: double sulphitation process, refining and grading of sugar. **Saccharin synthesis and uses** of sugar substitute. Ethanol: manufacture from molasses by fermentation. **Paper industry**; manufacture of paper: production of sulphite pulp and conversion to paper (bleaching, filling, sizing and calendaring) **12hrs**

UNIT III: Leather Chemistry: Introduction, structure of hides and skins. Outline of chief **processes used in leather manufacture**. Process before tannage (i) flaying (ii) curing (iii) soaking (iv) unhairing (v) liming (vi) fleshing (vii) deliming (viii) bating (ix) pickling. **Tanning Process** Methods of Tanning 1. Vegetable Tan 2. Chrome Tanning 3. Aldehyde Tanning. Finishing Process After Tannage Tannery effluent (i) primary Treatment (ii) Secondary Treatment (iii) Tertiary Treatment **12hrs**

UNIT IV: Agricultural Chemistry: Introduction to soil. **Soil Classification** & Survey; Properties of soil: Soil Texture; Soil water: Soil Temperature; Soil Colloids; Soil Minerals; Soil pH acidity and alkalinity; Buffering Soil; Soil Fertility, Soil Formation. **Fertilizers-Classification** of fertilizers, nitrogenous, fertilizers Phosphate fertilizers, Potash fertilizers, effect of fertilizers. **Insecticides:** DDT, BHC Herbicides: 2,4-D and 2,4,5-T **Fungicides:** Bordeaux mixture, Bio-gas, production and Manure. **12hrs**

UNIT IV: Chemistry in daily life: Synthetic food Additives, Preservatives, colourants and flavours Water: Domestic waste water treatment -Treatment by primary and secondary process Glass: composition, manufacture and uses Cement: Manufacture: wet and dry process, composition and setting of cement Soaps: Types of soaps, cleansing action of soaps, synthetic detergents. **12hrs**

RECOMMENDED TEXT BOOKS:

1. Jayshree Ghosh Fundamentals of Applied chemistry Sultan choud & co
2. R Gopalan P.S Subranian K. Rangarajan Elements of Analytical Chemistry Sultan & co

REFERENCE BOOKS:

1. Biswas. A.K., Frontiers in Applied Chemistry, Narosa publishing house, 1989
2. Vermain , O.P, Narula A.C, Applied Chemistry theory and books
3. G.T.Austin : shreve's Chemical Process Industries , 5th edition , Mc-Graw-Hill, 1984
4. B.A Yagodin (Ed). Agricultural Chemistry, 2 volumes, MIR publishers (Moscow), 1976.
5. Qutime of dairy technology- Sukarma De
6. Fundamentals of leather science- Woodroffe
7. BR Puri . IR Sharma, KV Kara Principles of Inorganic chemistry , Milstone publishers
8. M.K Jain, S V Sharma Modern organic chemistry Vishal publishers.

PERIODICALS:

1. Education in Chemistry
2. Current Science

WEBSITES & e-LEARNING SOURCES:

1. www.acs.org.
2. www.virtiab.com
3. Internet chemistry resources.

SEMESTER-VI

Title of the Course: **Core 12-Pharmaceutical Chemistry**

Teaching hours: 15 x 4= 60

Credits: **4**

Course Code: **CH15/6C/PHC**

L T P 3 1 0

Objectives:

1. To make the student understand important aspects of circulatory and respiratory systems
2. To realize the significance of clinical tests and hence health care management
3. To make the student to understand the physiological action of the drugs and its impacts

COURSE OUTLINE:

UNIT I: Introduction-Terminologies used-molecular pharmacology, pharmacodynamics, pharmacophore, metabolites, antimetabolites, bacteria virus fungi-Clinical chemistry-causes of common diseases and their treatment by drugs- Jaundice, piles, leprosy, epilepsy, typhoid, malaria, cholera, filarial, cancer-diagnostic tests for sugar, salt, cholesterol in blood and urine-First aid in emergencies. **12hrs**

UNIT II: **Antibiotics** Definition and mode of action. Structure and clinical uses of chloroamphenicol, penicillin, Streptomycin and tetracyclines. SAR and synthesis of chloramphenicol and penicillin only. **Hypoglycemic drugs and antineoplastic drugs.** **12hrs**

UNIT III: **Anaesthetics**-Definition-general and local classification- Gaseous-nitrous oxide, ethers halothane, trichloroethylene only - Intravenous anaesthetics. Thiopentone, sodium, methohexitone and propandiol, local anaesthetics-esters-cocaine-benzocaine, procaine, amides-cinchocaine-Tranquilizers, sedatives, hypnotics and psychodelic drugs (LSD, Hashish). **12hrs**

UNIT IV: **Analgesics, antipyretics and anti-inflammatory agents**-Classification, action and analgesics, narcotic and synthetic analgesics-pethidine and methadone. Salicylic acid and its derivatives, p-amino phenol derivatives, indolyl, aryl-acetic acid derivatives-clinical uses and adverse effects. **12hrs**

UNIT V: **Blood-Grouping,** composition, Rh factor, Blood pressure-hypertension and hypotension Function of plasma proteins. **Role of blood as oxygen carrier. Clotting mechanism-Hematological Agents-Anticoagulant and coagulant drugs-Heparin,** coumarine, citric acid, quinoxaline- Vitamin K, proteins, amino acids. **12hrs**

RECOMMENDED TEXT BOOKS:

1. Pharmaceutical chemistry, Dr.S.Lakshmi, 3rd edn. Sultan and sons
2. A text book of synthetic drugs O.D.Tyogi M.Yadav, 4th edn 1996 ANMOL publications Pvt.Ltd
3. A text book of pharmaceutical chemistry – Jayashree ghosh 1st edn Sultan Chand & sons.
4. Essentials of Bio chemistry –U.Satyanarayanan – Books & Allied (p) ltd
5. Organisation pharmaceutical chemistry by Harikisher singh etal
6. Synthetic drugs O.P Agarwal
7. Synthetic drugs- Gurdeep Chatwal's Medical laboratory, Ramnaik –Sood

REFERENCE BOOKS:

1. Text books of pharmaceutical chemistry Batlay and Drivers
2. Text book of pharmaceutical chemistry – Raviling
3. Bio chemistry – A L Lehninger
4. A text book of bio chemistry – S.Ambika
5. Medicinal chemistry – A Sutosh kar – New age

PERIODICALS:

1. Resonance- Journal of science education
2. Education in Chemistry
3. Indian academy of sciences-proceedings- Chemical Sciences
4. Current Science
5. Journal of Indian chemical education
6. Journal of American chemical society

WEBSITES & e-LEARNING SOURCES:

1. www.acs.org
2. www.virtlab.com
3. Internet chemistry resources

SEMESTER-VI

Title of the Course: **Elective 2- Polymer Chemistry**

Teaching hours: 15 x 5 = 75

Credits: **5**

Course Code: **CH15/6E/POC**

L T P 4 1 0

Objectives:

1. To introduce the concepts of polymers and polymerization
2. To provide comprehensive knowledge about the principles involved in the determination of molecular weight, glass-transition temperature, and synthesis of polymers using latest techniques
3. To introduce common organic and inorganic polymers and their applications

COURSE OUTLINE

UNIT I: **Classification of polymers**- natural polymers and synthetic polymers, synthetic polymers-Addition polymers, condensation polymers-**Mechanism** of polymerization-step reaction of polymerization, chain reaction polymerization-ionic, radical -one example each (individual mechanisms not included for cationic and anionic)-Stereospecific polymerization-Use of Ziegler Natta Catalysts-(Mechanism not needed)

15hrs

UNIT II: **Molecular weights** of polymers-number average and weight average methods. **End group analysis** (Number average). Cryoscopy and Ebullioscopy Weight average molecular weight-**Light scattering technique** Molecular weight distribution

15hrs

UNIT III: Structure and properties and reactions of polymers-**Primary and secondary bond forces** in polymers-coherence energy-structure property relationship 1.Mechanical property-Tensile Strength, flexural strength fatigue resistance, 2.Thermal stability-glass transition temperature 3.Flame resistance, 4.Chemical resistance 5.Degradability (Oxidative, Thermal, photochemical and Chemical, ultrasonic, high energy), cure reactions and polymer reactions.

15 hrs

UNIT IV: Important **industrial polymers preparation and applications** (mechanism not required) 1.Polyethylene, Poly propylene 2.Polyamides 3.Polyvinyl chloride, 4. polymethyl methacrylate. Polyesters, Polycarbonates, 5.Poly urethanes, 6.Phenol- Formaldehyde, Melamine-Formaldehyde, 7. Polysilane polysiloxanes

15 hrs

UNIT V: **Natural Polymers** and polymer material Rubbers, wool, silk, collagen, lignin, cellulose, casein General structure -**properties and application**. Composites- **Fibers**, laminated composites hybrid and nanocomposites structure, properties and applications. **Biomedical applications.**

15 hrs

RECOMMENDED TEXT BOOKS:

1. Introductory polymer chemistry – G.S.Mishra New Age 2007
2. Polymer science-V.R.Gowriker, N.V.Viswanathan and Jayadev sreedar New Age 2006
3. Principles of polymer science –P.Bhadur, N.V.Sastry, Alpha Science International Ltd 2005

REFERENCE BOOKS:

1. E.A.Collins, J.Bares and E.W.Billmeyer- Experiments in Polymer science, Wiley Inter science, Newyork 1973
2. Polymer science and Technology – Joel Fried 2nd edn., Prentice Hall PTR 2003
3. Polymer Chemistry an introduction M.P.Steven 3rd edn., Oxford Univ Press, USA 1999.

PERIODICALS:

1. Resonance- Journal of science education
2. Education in Chemistry
3. Indian academy of sciences-proceedings- Chemical Sciences
4. Current Science
5. Journal of Indian chemical education
6. Journal of American chemical society

WEBSITES & e-LEARNING SOURCES:

1. www.acs.org.
2. www.virtlab.com
3. Internet chemistry resources.

SEMESTER-VI

Title of the Course: **Elective 3-Nanochemistry**

Teaching hours: 15 x 5 =75

Credits: 5

Course Code: **CH15/6E/NAC**

L T P 4 1 0

Objectives:

1. To introduce and give an insight into the fascinating area of nanoscience and its development.
2. To learn the experimental techniques of nano scale synthesis, properties and characterization.
3. To understand the nano fabrications and their applications.

COURSE OUTLINE

- UNIT I:** **Introduction to Nanochemistry:** origin, nano in ages, development of nanotechnology. Nanomaterials, nano compositions. Background to nanoscience and nanotechnology- scientific revolutions- nanosized effects- surface to volume ratio-atomic structure-molecules & phases- **Nanotechnology- importance of the nano scale materials and their devices.** 15 hrs
- UNIT II:** **Experimental Techniques Preparation:** **Bottomup and top down methods.** **Mechanical methods:** grinding material-ball ratio-medium for grinding Physical methods: vapour deposition and different types of epitaxial growth techniques-pulsed laser deposition, magnetron sputtering-micro lithography. **Chemical method:** sol-gel technique-**solvothetical methods**-control of grain size-co-precipitation hydrolysis-sonochemical method combustion technique-colloidal precipitation template process. 15 hrs
- UNIT III:** **Characterization:** **Basic principles and applications of UV-Vis, FTIR, FT-Raman, Photoluminescence, Light scattering methods.** X-ray techniques: X-ray powder diffraction-determination of crystallite size- Scherrer equation and its limitations, qualitative determination of phases, structure analysis. 15 hrs
- UNIT IV:** **Nano Scale Properties and Materials:** Size dependent variation in physical and chemical, optical properties, mechanical-micro hardness fracture toughness. magnetic properties -concepts of dia, para,ferro and ferri magnetism, electronic transport. Electrical properties-electrical conductivity. **Nanomaterials-Fillers, fibres, Wires, Fluids, Tubes- CNT, grapheme, Nanoclays.** 15 hrs

UNIT V: **Applications:** Ceramics and Composites. Molecular electronics, sensors, catalysts, molecular recognition, social implications of Nano science and technology. Nano materials interface with biological systems; environmental and health aspects.

15hrs

RECOMMENDED TEXTBOOKS:

1. Nanoscience and Nanotechnology: Fundamentals to Frontiers, M.S. Ramachandra Rao, Shubra Singh Wiley, India, First edition 2013.
2. Nano: The essentials, T. Pradeep Tata McGraw-Hill Publishing Company Limited, 2007.
3. Nanoscale materials in chemistry, Kenneth, J. Klabunde willey Interscience, 2001
4. Nano Materials: B.Viswanathan, Narosa Publishing House, 2000.
5. Nanoscale materials in chemistry, Kenneth, J. Klabunde willey Interscience, 2001
6. Nano: The essentials, T. Pradeep Tata McGraw-Hill Publishing Company Limited, 2007.

REFERENCE BOOKS

1. Nano and Microelectro mechanical systems: Fundamentals of Nano and microengmeering Sergy Lyshevski, Vallabh publications 2/e 2005.
2. Nano chemistry: A chemical approach to nanomaterials Geoffrey A.Ozin, Andre C. Arsenault RSC Publishing, London.2005
3. Nanostructure & Nanomaterials: Synthesis, properties & Applications G. Cao, Imperial College Press, 2004.
4. Nanomaterials, Nanotechnologoes and Design: An introduction for engineers and Architects Micheal F. Ashby. P.J. Ferria. D.L. Schodek. Palgrave Macmilan 2005.

PERIODICALS:

1. Current Science
2. Journal of Nanotechnology
3. Journal of Nanomaterials

WEBSITES &e-learning sources

5. www.nano.org
6. www.virtlab.com
7. Internet chemistry resources

SEMESTERS – V & VI

MAJOR PRACTICAL

Title of the Paper: **Core 13- Practical 3- Organic Chemistry**

Teaching hours: 30 x 4 = 120

Credits: **3**

Course Code: **CH15/6C/PR3**

1. **Determination of Boiling point of liquids**
- * 2. **Estimation of Aniline**
- *3. **Preparation** based on the following reactions
 - (a) Oxidation
 - (b) Nitration
 - (c) Bromination
 - (d) Esterification
 - (e) Hydrolysis
 - (f) Diazotisation & Coupling
4. **Analysis of organic compounds** with one functional group & characterization with one derivative –Aromatic Aldehyde, Aromatic / Aliphatic Ketone ,Aromatic/ Aliphatic, Mono, Di, Unsaturated and saturated Carboxylic acid, Primary, secondary & Tertiary Amine, Simple Phenol, Resorcinol & β - Naphthol, Aromatic Ester, Aromatic mono amide and Diamide, Anilide, Mono Nitro and poly nitro compound and Carbohydrate (Monosaccharide).

The students will report on the following

- Aromatic/Aliphatic
- Saturated /Unsaturated
- Elements present [Nitrogen, Sulphur and halogens]
- Functional group
- Solid derivative

.*For internal assessment only.

END SEMESTER PRACTICAL EXAMINATION

III B.Sc SEMESTERS-V & VI

QUESTION BANK

Title of the Paper: Core 13- Practical-3 Organic Chemistry

Time: 3hrs

Course Code: CH15/6C/PR3

Max. Marks: 60

- (1) Systematically carry out suitable test on the given organic substance and report on the following
- Aliphatic/ Aromatic
 - Saturated/ Unsaturated
 - Elements present or absent-(Nitrogen, Sulphur and Halogens)
 - Functional group

Confirm the above by a suitable colour reaction and prepare atleast one solid derivative and submit it for evaluation

Substances to be given-Benzaldehyde, Acetophenone, 2- Butanone , Glucose, Ethyl benzoate, Cinnamic acid, Succinic acid , Phthalic acid, Benzoic acid , Phenol ,Resorcinol, β - Naphthol, Nitrobenzene, Aniline, N,N- Dimethylaniline, Acetanilide, Benzamide, Urea

- (2) Determine the boiling point of given liquid –Water, butan-2-one, butan-1-ol, Propan-2-ol

Scheme of valuation

Total marks = 100 (CA 40 and End Semester 60)

Organic analysis- 35 marks

Physical constant-15 marks

Record -10marks

60 marks

- Organic analysis-(a)Aliphatic/ Aromatic-4 marks (2x2)
(b)Saturated/ Unsaturated-4 marks (2x2)
(c)Elements present or absent-6 marks (3x2)
(d)Preliminary reaction -10 marks
(e) functional group-6 marks
(f) Derivative-5 marks

Physical constant- $\pm 2^\circ\text{C}$ - 15 marks

$\pm 3^\circ\text{C}$ - 13 marks

$\pm 4^\circ\text{C}$ - 10 marks

$> 4^\circ\text{C}$ - 6 marks

SEMESTERS-V& VI

MAJOR PRACTICALS

Title of the Course: **Core 14: Practical 4- Gravimetric Analysis**

Teaching hours: 30 x 4 = 120

Credits: **3**

Course Code: **CH15/6C/PR4**

1. Estimation of Lead as Lead chromate.
2. Estimation of Zinc as Zinc Oxinate
3. Estimation of Barium as Barium chromate.
4. Estimation of Barium as Barium sulphate.
5. Estimation of calcium as calcium oxalate monohydrate.
- * 6. Estimation of Nickel as Nickel dimethyl glyoximate complex.
- * 7. Estimation of copper as cuprous thiocyanate.
- * 8. Estimation of chloride as silver chloride.

* INTERNAL ASSESSMENT ONLY

END SEMESTER PRACTICAL EXAMINATION

III B.Sc Chemistry- SEMESTERS – V & VI

QUESTION BANK

Title of the Course: Core 14: Practical 4- Gravimetric Analysis

Course Code: CH15/6C/PR4

**Time: 3 hrs
Max. Marks: 60**

1. Estimate the amount of barium present in the whole of the given solution. You are provided with 5% K_2CrO_4 solution.
2. Estimate the amount of sulphate present in the whole of the given solution. You are provided with 4% $BaCl_2$ solution.
3. Estimate the amount of barium present in the whole of the given solution. You are provided with 2N H_2SO_4 solution.
4. Estimate the amount of Lead present in the whole of the given solution. You are provided with 5% K_2CrO_4 solution.
5. Estimate the amount of Zinc present in the whole of the given solution. You are provided with 1% Oxine solution.

SCHEME OF VALUATION

Continuous Assessment –	40 marks
End Semester –	60 marks (50 + 10 marks for Record)
Total	100 marks
Error up to 2%	- 50 marks
2.1 - 3%	- 45 marks
3.1-4%	- 35 marks
>4 %	- 20 marks

SEMESTERS –V & VI

MAJOR PRACTICALS

Title of the Course: **Core 15- Practical 5-Physical Chemistry**

Teaching hours: 30 x 4 =120

Credits: **3**

Course Code: **CH15/6C/PR5**

1. **KINETICS**

Determination of rate constant

- i. I order kinetics - Acid catalysed hydrolysis of an ester
- ii. II order kinetics - Reaction between potassium perdisulphate and potassium iodide
- iii. Zero order kinetics - Acid catalysed iodination of acetone.

2. **Determination of K_f and Molecular Weight** by Rast's Macro method

3. **PHASE RULE**

- i. **CST** of Phenol-water system
- ii. **Effect of electrolyte on CST** of phenol –water system and determination of concentration of the electrolyte.
- iii. **Determination of transition temperature** of hydrated salts- Sodium acetate, sodium thiosulphate, strontium chloride (any one)

4. **ELECTROCHEMISTRY**

Conductivity

- i. Determination of cell constant
- ii. Equivalent conductance of strong and weak electrolyte solutions of atleast five different concentrations and to determine Λ^∞ for strong electrolyte
- iii. Strong acid Vs Strong base - HCl against NaOH
- iv. *Weak acid Vs strong base - CH₃COOH against NaOH
- v. *Dissociation constant of a weak acid

5. ***DISTRIBUTION LAW**

Determination of

- i. *Distribution coefficient of iodine between water and CCl_4
- ii. *Equilibrium constant of the equilibrium
$$\text{KI} + \text{I}_2 = \text{KI}_3$$
- iii. *Association factor of benzoic acid in benzene.

6. ***POTENTIOMETRY**

*Acid base titration – HCl against NaOH

* For Internal Assessment only.

END SEMESTER PRACTICAL EXAMINATION

III B.Sc. CHEMISTRY – SEMESTERS V & VI

QUESTION BANK

Title of the Course: Core 15- Practical 5-Physical Chemistry

Time: 3hrs

Course Code: CH15/6C/PR5

Max. Marks: 60

1. Determine the rate constant for the acid catalysed hydrolysis of the given ester with the given acid solution at room temperature.
2. Determine the rate constant of the reaction between potassium iodide and potassium persulphate at room temperature.
3. Find out the rate constant for the iodination of acetone at room temperature, using the given acid solution as catalyst.
4. Determine the molecular weight of the given solute. You are provided with a suitable solvent whose K_r value is-----
5. Determine the K_r of the given solvent. You are provided with a suitable solute whose molecular weight is -----
6. Find out the concentration of the given sodium chloride solution. You are provided with pure phenol and solutions of sodium chloride.
7. Determine the transition temperature of the given hydrated salt, by thermometric method.
8. Find out the cell constant of the given conductivity cell using 0.1M and 0.01M potassium chloride solutions, whose specific conductivities are given. Determine the equivalent conductance of the given two solutions of known concentrations.
9. Determine the strength of the given hydrochloric acid by conductometric titration. You are provided with sodium hydroxide of known strength

SCHEME OF VALUATION

External Marks (60) + Internal Marks (40) = 100 Marks

Record	-	10 marks
Manipulation	-	15 marks
Experiment	-	35 marks

60 marks

1.	KINETICS		Marks
	Below a factor of 10	-	35
	By a factor of 10	-	25
	Above	-	10
2.	MOLECULAR WEIGHT & K_f		Marks
	Error upto 10%	-	35
	10 – 20%	-	30
	Upto 30%	-	20
	> 30%	-	10
3.	PHENOL – NaCl		Marks
	Error upto 10%	-	35
	10 – 20%	-	30
	Upto 30%	-	20
	> 30%	-	10
4.	TRANSITION TEMPERATURE		
	Error upto 2°C	-	35
	Upto 4°C	-	30
	Upto 6°C	-	20
	> 6°C	-	10
5.	CONDUCTIVITY		
			1/a Λ_1 Λ_2
	Error upto 10%	-	15+10+10
	Upto 15%	-	10+8+8
	> 15%	-	5+4+4
6.	CONDUCTOMETRIC TITRATION		
	Error upto 10%	-	35
	10 – 20%	-	30
	Upto 30%	-	20
	> 30	-	10

SEMESTER- I
(For I B.Sc N&D)

Title of the course: **Allied Chemistry-General Chemistry**

Teaching hours: 15 x 4 = 60
Course Code: **CH15/1A/NGC**

Credits: 4
L T P 3 1 0

Objectives:

1. To introduce the basic aspects of structure and bonding in Organic and Inorganic compounds and simple separation techniques
2. To create an awareness on organic compounds used as dyes.
3. To impart basic knowledge on aromatic hydrocarbons and electrophilic substitution reactions

COURSE OUTLINE

UNIT I: **Aromatic Compounds:** Structure of benzene (Structural elucidation is not necessary). Electrophilic substitution in benzene, Mechanism in nitration, sulphonation, halogenation, alkylation and acylation. **Naphthalene and Anthracene-**isolation, **preparation, properties and uses.** Structural elucidation of naphthalene. **12 hrs**

UNIT II: **Stereo isomerism:** **Optical isomerism,** elements of symmetry, cause of optical activity, Lactic acid and tartaric acid, racemisation, resolution, **R&S configuration,** Geometrical isomerism of maleic and fumaric acid, **keto-enol tautomerism.** **12hrs**

UNIT III: **Biochemical techniques and Dyes:** **Chromatography-**Principle and application of column, paper, thin layer and Ion exchange chromatography. Electrophoresis-Principles, apparatus and application. Ultra centrifugation-principle. **Dyes-****Classification-** Triphenyl methane dyes, Malachite green, p-rosaniline, azodyes- methyl orange, aniline yellow, food colours. **12hrs**

UNIT IV: **Chemical bonding:** **Molecular orbital theory,** bonding, antibonding and non bonding orbitals. Molecular orbital configuration of H₂, N₂, O₂ and Fluorine. Bond order, Diamagnetism and paramagnetism. **Hydrides-** classification, preparation and properties. Diborane, Sodium borohydride and Borazole-preparation, **properties and structure.** **12 hrs**

UNIT V: **Coordination Chemistry:** Nomenclature, Coordination Compounds, shapes of d-orbitals, Theories of Co-ordination complexes- **Werner's theory, Sidgwicks theory** and **Pauling's theory.** Theory and structure of Chelation, Haemoglobin and Chlorophyll. **Application of complexes-**Qualitative and Quantitative analysis. **12hrs**

RECOMMENDED TEXT BOOK:

1. Allied Chemistry- Gopalan and Sundaram, 3rd edn., Sultan Chand and Sons.

REFERENCE BOOKS:

1. Organic Chemistry- P.L.Soni, 28th edn, Sultan Chand and Sons (1999)
2. Principles of Inorganic Chemistry- B.R.Puri, L.R.Sharma & K.C.Kalia 28th edn, Vallabh publications (2003).
3. Text book of Reaction mechanism in Organic Chemistry- Peter sykes, Orient Longman.

WEBSITES & e-LEARNING SOURCES:

1. www.acs.org
2. www.virtlab.com
3. Internet Chemistry resources
4. MATLAB

SEMESTER-II
(For I B.Sc N&D)

Title of the course: **Allied Chemistry- Bio organic Chemistry**

Teaching hours: 15 x 4 = 60

Credits: 4

Course Code: **CH15/2A/NBC**

L T P 3 1 0

Objectives:

1. To impart basic knowledge on carbohydrates, proteins, lipids and vitamins.
2. To introduce mechanism of carbohydrate and lipid metabolisms.
3. To understand the role of amino acids and enzymes.

COURSE OUTLINE

UNIT I: **Carbohydrates and its metabolism:** Classification-preparation and properties of **monosachharides**- glucose and fructose, **discaccharides**-Sucrose, Polysaccharides-starch and cellulose. Differences between starch and cellulose. Purine and Pyrimidine-types of bases, nucleosides, nucleotides and polynucleotide. **DNA**-Structure, replication, **RNA**-Types of RNA, structure and functions, genetic code. **Blood composition**-Plasma proteins, Blood coagulation. **12hrs**

UNIT II: **Digestion and absorption of carbohydrates:** **Metabolism of glucose**, glycolysis, TCA cycle, glycogenolysis, glycogenesis and gluconeogenesis. **Enzymes**-Definition, classification, nomenclature, specificity, isoenzymes, factors affecting enzyme activity-substrate concentration, pH and temperature, **Michaelis-Menten equation**, enzyme inhibitors- competitive, non-competitive and uncompetitive. **12hrs**

UNIT-III: **Hormones:** Definition, classification, **mechanism of action of hormones**-General functions. Posterior pituitary hormones-Oxytocin and anti-diuretic- Structure and function. Chemotherapy- **Preparations, uses and mode of action of sulpha drugs**, structure and **use of penicillin and chloromycetin**. **12hrs**

UNIT IV: **Vitamins:** Definition, **Classification**-Water soluble vitamins B₁, B₂, B₃, B₆ and B₁₂ and Vitamin C. Fat soluble vitamins-A, D, E, K-Occurrence, deficiency diseases, biochemical roles and daily requirements. **Steriods**- Cholesterol, physical properties and biological function only. **12hrs**

UNIT-V: **Lipids and its metabolism:** **Classification of lipids**-saponifiable and non saponifiable lipids, phospholipids- different types, their constituents, functions. Fatty acids-saturated - stearic and palmitic, formulae, unsaturated - oleic acid and essential fatty acids. **Characterisation by iodine value**, RM Value, acid number, saponification value (Definitions only). **Amino acids**- Classification and structures, amphoteric nature,

isoelectric point, peptide bond. **Proteins**-Classification, Primary structure, protein digestion, transamination, oxidative deamination and urea cycle- inter -relationship of protein. **12hrs**

RECOMMENDED TEXT BOOK:

1. Text book of BioChemistry – Ambika Shanmugam.

REFERENCE BOOKS:

1. Text book of natural products by chatwal –vol I Himalayan Publication.
2. Text book of BioChemistry - Agarwal, Goel Publications.
3. Fundamentals of BioChemistry – J. L. Jain, Sultan Chand and Sons.
4. Text book of Pharmaceutical Chemistry- Jayashree Ghosh

WEBSITES & e-LEARNING SOURCES:

1. www.acs.org
2. www.virtlab.com
3. Internet Chemistry resources
4. MATLAB

SEMESTER- III

(For II B.Sc AZB)

Title of the course: Allied Chemistry-General Chemistry

Teaching hours: 15 x 4 = 60

Credits: 4

Course Code: CH15/3A/ZGC

L T P 3 1 0

Objectives:

1. To introduce the basic aspects of structure and bonding in Organic and Inorganic compounds and simple separation techniques.
2. To create an awareness on Organic compounds used as dyes.
3. To impart basic knowledge on aromatic hydrocarbons and electrophilic substitution reactions.

COURSE OUTLINE

UNIT I: **Aromatic Compounds:** Structure of benzene (Structural elucidation is not necessary). Electrophilic substitution in benzene, Mechanism in nitration, sulphonation, halogenation, alkylation and acylation. Naphthalene and Anthracene-isolation, preparation, properties and uses. Structural elucidation of naphthalene. **12 hrs**

UNIT II: **Stereoisomerism:** Optical isomerism, elements of symmetry, cause of optical activity, Lactic acid and tartaric acid, racemisation, resolution, R&S configuration, Geometrical isomerism of maleic and fumaric acid, keto-enol tautomerism. **12 hrs**

UNIT III: **Biochemical techniques and Dyes:** Chromatography – Principle and application of column, paper, thin layer and Ion exchange chromatography. Electrophoresis - Principles, apparatus and application. Ultra centrifugation-principle. Dyes- Classification- Triphenyl methane dyes, Malachite green, p-rosaniline, azodyes- methyl orange, aniline yellow, food colours. **12hrs**

UNIT IV: **Chemical bonding:** Molecular orbital theory, bonding, antibonding and non bonding orbitals. Molecular orbital configuration of H₂, N₂, O₂ and Fluorine. Bond order, Diamagnetism and paramagnetism. Hydrides- classification, preparation and properties. Diborane, Sodium borohydride and Borazole-preparation, properties and structure. **12hrs**

UNIT V: **Coordination Chemistry-** Nomenclature, Coordination Compounds, shapes of d-orbitals, Theories of Co-ordination complexes- **Werner's theory, Sidgwicks theory** and **Pauling's theory.** Theory and structure of Chelation, Haemoglobin, and Chlorophyll. **Application of complexes-**Qualitative and Quantitative analysis.

12 hrs

RECOMMENDED TEXT BOOK:

1. Allied Chemistry- Gopalan and Sundaram, III edn., Sultan Chand and Sons.

REFERENCE BOOKS:

1. Organic Chemistry- P.L.Soni, 28th edn, Sultan Chand and Sons (1999)
2. Principles of Inorganic Chemistry- B.R.Puri, L.R.Sharma & K.C.Kalia, 28th edn, Vallabh publications (2003).
3. Text book of Reaction mechanism in Organic Chemistry- Peter sykes, Orient Longman.

WEBSITES & e-LEARNING SOURCES:

1. www.acs.org
2. www.virtlab.com
3. Internet Chemistry resources
4. MATLAB

SEMESTER-IV

(For II B.Sc AZB)

Title of the course: **Allied Chemistry- Bio organic Chemistry**

Teaching hours: 15 x 4 = 60

Credits: 4

Course Code: **CH15/4A/ZBC**

L T P 3 1 0

Objectives:

1. To impart basic knowledge on carbohydrates, proteins, lipids and vitamins.
2. To introduce mechanism of carbohydrate and lipid metabolisms.
3. To understand the role of amino acids and enzymes.

COURSE OUTLINE

- UNIT I:** **Carbohydrates and its metabolism:** Classification – preparation and properties of **monosachharides**- glucose and fructose, **discaccharides**-Sucrose, **Polysaccharides**-starch and cellulose. Differences between starch and cellulose. Purine and Pyrimidine- types of bases, nucleosides, nucleotides and polynucleotide. **DNA**-Structure, replication, **RNA**-Types of RNA, structure and functions, genetic code. Blood composition-Plasma proteins, Blood coagulation. **12 hrs**
- UNIT II:** **Digestion and absorption of carbohydrates:** **Metabolism of glucose**, glycolysis, TCA cycle, glycogenolysis, glycogenesis and gluconeogenesis. **Enzymes**-Definition, classification, nomenclature, specificity, isoenzymes, factors affecting enzyme activity-substrate concentration, pH and temperature, **Michaelis-Menten equation**, enzyme inhibitors- competitive, non-competitive and uncompetitive. **12 hrs**
- UNIT-III:** **Hormones:** Definition, classification, **mechanism of action of hormones**-General functions. Posterior pituitary hormones-Oxytocin and anti-diuretic- Structure and function. **Chemotherapy- Preparations, uses and mode of action of sulph drugs**, structure and use of penicillin and chloromycetin. **12hrs**
- UNIT IV:** **Vitamins:** Definition, **Classification-Water soluble vitamins** B₁,B₂,B₃,B₆ and B₁₂ and Vitamin C. Fat soluble vitamins-A,D,E,K-Occurrence, deficiency diseases, biochemical roles and daily requirements. **Steriods- Cholesterol**, physical properties and biological function only. **12hrs**
- UNIT-V:** **Lipids and its metabolism:** **Classification of lipids**-saponifiable and non saponifiable lipids, phospholipids- different types, their consitutents, functions. Fatty acids- saturated - stearic and palmitic, formulae, unsaturated - oleic acid and essential fatty acids. **Characterisation by iodine value, RM Value, acid number, saponification value (Definitions only).** **Amino acids- C**lassification and structures,

amphoteric nature, isoelectric point, peptide bond. **Proteins**-Classification, Primary structure, protein digestion, transamination, oxidative deamination and urea cycle-inter-relationship of protein. **12hrs**

RECOMMENDED TEXT BOOK:

1. Text book of Bio chemistry – Ambika Shanmugam.

REFERENCE BOOKS:

1. Text book of natural products by Chatwal –vol I Himalayan Publication.
2. Text book of BioChemistry - Agarwal, Goel Publications.
3. Fundamentals of BioChemistry – J. L. Jain, Sultan Chand and Sons.
4. Text book of Pharmaceutical Chemistry- Jayashree Ghosh

WEBSITES & e-LEARNING SOURCES:

1. www.acs.org
2. www.virtlab.com
3. Internet Chemistry resources
4. MATLAB

SEMESTER- III
(For II B.Sc. Physics)

Title of the course: **Allied Chemistry-General Chemistry 1**

Teaching hours: 15 x 4 = 60

Credits: 4

Course Code: **CH15/3A/PGCI**

L T P 3 1 0

Objectives:

1. To introduce the concepts of qualitative and quantitative analysis and separation techniques
2. To provide the basic concepts in organic chemistry and an exposure to the applications of coordination chemistry
3. To provide a basic knowledge in photochemistry and electrochemistry

COURSE OUTLINE

UNIT I: **Analytical Chemistry-** Introduction to Quantitative and Qualitative Analysis- Principle of volumetric analysis – separation techniques-extraction-distillation-crystallization-chromatographic separations-Principles and application of column, paper, thin layer, gas liquid and ion-exchange. **12hrs**

UNIT II: **Coordination chemistry:** Definition of terms, classification of ligands, Nomenclature of coordination compounds. Theories of co-ordination complexes- Werners theory, Sidgwick's theory and Pauling theory. Chelation, Biological role of Haemoglobin and Chlorophyll, Application of complexes- Qualitative and Quantitative analysis. **12hrs**

UNIT III: Inductive, electromeric, mesomeric, hyperconjugative and steric effects-influence of effects in properties of compounds-Acidity and Basicity. Aromaticity-Huckles Rule concept of delocalization, Resonance energy-Benzene, Electrophilic substitution in benzene- Mechanism of nitration, halogenation, alkylation, acylation and sulphonation. **12hrs**

UNIT IV: **Photochemistry-** Beer-Lambert's law, Grothurs-Draper's law, Stark Einstein's law of photochemical equivalence. Quantum yield- Examples with hydrogen and chlorine reaction. Photosynthesis, Jablonski diagram-Radiative process- Fluorescence, Phosphorescence, non-radiaative process-Internal conversion and Intersystem crossing, Chemiluminescence, Photosensitization-only definition with examples. **12hrs**

UNIT V: **Electrochemistry:** Definition of specific, equivalent and molar conductance and their determination, effect of dilution on conductance, Ostwald's dilution law,

Kohlrausch's law and its applications Galvanic cells, Emf, Standard cell-Weston cadmium cell. Standard hydrogen electrode and calomel electrode in e.m.f. measurements Standard electrode potentials and its applications, Henderson equation, applications of pH and buffer in biological systems and Industries.

12hrs

RECOMMENDED TEXT BOOKS:

1. Allied Chemistry- Gopalan and Sundaram, III edn., Sultan Chand and Sons.

REFERENCE BOOKS:

1. Organic chemistry, P.L.Soni 28th edn., Sultan Chand and Sons 1999.
2. Principles of Inorganic chemistry, B.R.Puri, L.R.Sharma & K.C.Kalia 28th edn., Vallabh publications 2003.
3. Principles of Physical chemistry – Puri & Sharma 41st edn., Vishal Publishing Co. 2004.
4. Elements of Analytical Chemistry –R.Gopalan,P.S.Subramanian and K.Rengarajan Sultan Chand and Sons.

WEBSITES & e-LEARNING SOURCES:

1. www.acs.org
2. www.virtlab.com
3. Internet chemistry resources

SEMESTER-IV
(For II B.Sc Physics)

Title of the course: **Allied Chemistry- General Chemistry 2**

Teaching hours: 15 x 4 = 60

Credits: **4**

Course Code: **CH15/4A/ PGC2**

L T P 3 1 0

Objectives:

1. To expose the students to the industrial applications of silicones and fertilizers
2. To introduce the concepts of stereo isomerism in organic chemistry and basic knowledge about chemotherapy, vitamins and hormones
3. To introduce the concept of phase rule and phase diagram

COURSE OUTLINE

- UNIT I:** **Industrial Chemistry:** Silicones-synthesis, properties and uses. Fuel gases- Natural gas, water gas, semiwater gas, carbureted water gas, producer gas. Inter halogen compounds ICl, BrF₃, IF₅, IF₇- preparation, properties, hybridization and structure, Fertilisers-Urea, Super phosphate of lime, Triple super phosphate. **12hrs**
- UNIT II:** **Metallurgy:** General methods of extraction of metals. Types of ores, methods of ore dressing, types of furnaces, reduction methods, types of refining, Van Arkel, Zone refining, Extraction of copper, uranium and Thorium **12hrs**
- UNIT III:** **Chemotherapy:** Preparation and uses of sulphapyridine, sulphadiazine, sulphaguanidine, sulpha methazine, mode of action of sulpha drugs. Antibiotics- penicillin, chloromycetin and Streptomycin-structure and uses only (elucidation not necessary). Analgesics – Paracetamol and aspirin, Vitamins- A, B₁, B₂, C and D- sources and deficiency diseases of vitamins. Hormones- Peptide hormones- Oxytocin, Protein hormones- Insulin (biological functions only). **12hrs**
- UNIT IV:** Geometrical isomerism of maleic and fumaric acid, cis, trans, E & Z nomenclature of geometrical isomers, Optical isomerism, elements of symmetry, cause of optical activity, Specific rotation-Determination of Specific rotation, Lactic acid and tartaric acid, Resolution, Racemisation and Walden inversion, Configuration-D&L, R&S. **12hrs**
- UNIT V:** **Phase rule:** Definition of terms in phase rule, Gibbs phase rule (no derivation) one component systems-water and sulphur systems, Reduced phase rule-two component system, Simple eutectic system. Eg. Lead-silver & Potassium iodide-water system **12hrs**

RECOMMENDED BOOKS:

1. Allied Chemistry- Gopalan and Sundaram, III edn.,Sultan Chand and Sons.

REFERENCE BOOKS:

1. Organic chemistry P.L.Soni 28th edn Sultan Chand and Sons 1999
2. Principles of Inorganic chemistry B.R.Puri, L.R.Sharma & K.C.Kalia 28th edn, Vallabh publications 2003
3. Principles of physical chemistry – Puri & Sharma 41st edn Vishal Publishing Co. 2004

WEBSITES & e-LEARNING SOURCES:

1. www.acs.org
2. www.virtlab.com

ALLIED CHEMISTRY PRACTICAL

I YEAR- SEMESTER I & II
(I B.Sc. N&D)
&
II YEAR- SEMESTER III & IV
(for II B.Sc. AZB & II B.Sc. PHYSICS)

Title of the Course: **Allied Chemistry Practical-Volumetric & Organic Analysis**

Teaching hours: 30 x 2= 60

Credits: **2**

Course Code: **CH15/2A/PR & CH15/4A/PRA**

VOLUMETRIC ANALYSIS

Acidimetry- Alkalimetry

1. Estimation of Borax using Methyl orange as indicator.
2. Estimation of Oxalic acid using Phenolphthalein as indicator.

Permanganimetry

3. Estimation of ferrous ammonium sulphate.
4. Estimation of oxalic acid.

Complexometry

5. Estimation of Magnesium/Zinc using EDTA

Dichrometry

- 6.*Estimation of ferrous ion using diphenyl amine as internal indicator.

Iodometry

- 7.*Estimation of copper sulphate using standard potassium dichromate.

***For internal assessment only**

ORGANIC ANALYSIS

Systematic Analysis of Organic Compounds with one functional group- Aromatic Aldehyde, Mono and dihydric phenols, β -Naphthol, Aromatic monocarboxylic acid, Aliphatic and Aromatic dicarboxylic acids, Carbohydrates-monosaccharide, Aliphatic diamide and Primary Aromatic Amine

1. Detection of elements- Nitrogen, Sulphur and Halogens.
2. Identification of Aliphatic or Aromatic compounds.
3. Identification of Saturated or Unsaturated compounds.
4. Identification of Functional group

END SEMESTER PRACTICAL EXAMINATIONS
(For I B.Sc N&D, II B.Sc. AZB & II B.Sc. PHYSICS)
SEMESTER II/IV

QUESTION BANK

Title of the Paper: Allied Chemistry Practical-Volumetric & Organic Analysis

Paper Code: CH15/2A/PRA & CH15/4A/PRA

Time: 3hrs

Max. Marks: 60

1. Analyze volumetrically the amount of sodium carbonate present in the whole of the given solution. You are provided with a standard solution of sodium hydroxide and a link solution of hydrochloric acid.
2. Estimate the amount of Borax present in the whole of the given solution. You are given a standard sodium carbonate solution and a link hydrochloric acid solution.
3. Estimate the amount of hydrochloric acid present in the whole of the given solution. You are provided with a standard solution of sulphuric acid and a link solution of sodium hydroxide.
4. Estimate the amount of oxalic acid present in the whole of the given solution. You are provided with a standard solution of hydrochloric acid and a link solution of sodium hydroxide.
5. Estimate the amount of Ferrous sulphate present in the whole of the given solution. You are provided with a standard solution of ferrous ammonium sulphate and a link solution of potassium permanganate.
6. Estimate the amount of Ferrous ammonium sulphate present in the whole of the given solution. You are provided with a standard solution of oxalic acid and a link solution of potassium permanganate.
7. Estimate the amount of oxalic acid present in the whole of the given solution. You are provided with a standard solution of ferrous sulphate and a link solution of potassium permanganate.
8. Estimate the amount of Magnesium sulphate present in the whole of the given solution. You are provided with a standard solution of Magnesium sulphate and a link solution of EDTA.

SCHEME OF VALUATION

Total Marks - 100

External Marks (60) + Internal Marks (40) = 100 Marks

Record	- 10
Volumetric analysis	- 25
Organic Analysis	- 25

Total Marks	60

Scheme of Valuation: Volumetric Analysis

Error upto 2% - 25 marks

Error upto 2.1 - 3% - 20 marks

Error upto 3.1 - 4% - 15 marks

Error > 4% - 10 marks

Arithmetic error - Reduce 1 mark.

Wrong calculation - Reduce 20% of the marks

No calculation - Reduce 40% of the marks

Analyse the given organic substance and report for the following.

- Aromatic or Aliphatic
- Saturated or Unsaturated
- Presence or Absence of the elements - Nitrogen, Sulphur and Halogens
- Functional group present

List of substances given for Analysis

1. Benzaldehyde
2. Phenol
3. β -Naphthol
4. Resorcinol
5. Glucose
6. Benzoic acid
7. Phthalic acid
8. Succinic acid
9. Aniline
10. Urea

Scheme of Valuation: Organic Analysis: 25 Marks

Elements	(3 x 2)	6 Marks
Aliphatic or Aromatic	(2 tests x 2)	4 Marks
Saturated or Unsaturated	(2 tests x 2)	4 Marks
Preliminary reactions		6 Marks
Functional group tests		5 Marks

PART-IV NON MAJOR ELECTIVE

SEMESTER-I

(Common to I Year B.A/B.Sc./B.Com)

Title of the Course: Non Major Elective 1 - Cosmetology

Teaching hours: 15 x 2 = 30

Credits: 2

Course Code: CH15/IN/COS

L T P 2 0 0

Objectives:

1. To improve one's own personal appearance and develop self confidence.
2. To impart skill in cosmetology.
3. To acquire knowledge about cosmetic science.

COURSE OUTLINE:

UNIT I: Self analysis, grooming, professional behavior, skin care, theory of massage, facials, facial manipulation, make-up- purpose and types- corrective make-up using optical illusions. **10hrs**

UNIT II: Hair oils and hair creams, hair removers- temporary removal of hair, depilation process and epilation process. Permanent removal of hair. Hand lotions and creams- Preparation and applications. Hazards of cosmetics and quality control. **10hrs**

UNIT III: Face creams: Types of face creams-cold, vanishing, cleaning and bleaching creams- preparation and application. Hand lotions and creams-simple method of preparation. Facial, manicure, pedicure, waxing, make-up & hair style. **10hrs**

RECOMMENDED TEXT BOOKS:

1. Text book of Cosmetology by Mary Haely Eastern Economy edition.
2. The complete book of beauty care- Aruna Anand.

WEBSITES & e-LEARNING SOURCES:

1. www.acs.org
2. www.virtlab.com
3. Internet Chemistry resources
4. MATLAB

PART IV – NON MAJOR ELECTIVE

SEMESTER – II

(Common to I B.A./B.Sc./B.Com.)

Title of the Course: **Non Major Elective 2- Dyeing and Printing of Textiles**

Teaching hours: 15 x 2 = 30

Credits: **2**

Course Code: **CH15/2N/DPT**

L T P 2 0 0

Objective:

1. To have a knowledge about the types and characteristics of dyes
2. To impart skill in dyeing and printing of textiles
3. To inculcate entrepreneur skills in textile dyeing

COURSE OUTLINE

UNIT I: **Dyes:** Classification-Main types of dyes, and their characteristics-**Natural dyes** and **synthetic dyes**, Natural dyes based on plant origin. Synthetic dyes-cationic, anionic, direct dyes, mordant dyes, vat dyes and sulphur dyes. **10hrs**

UNIT II: Basic methods of **dyeing**-stock dyeing, Top dyeing, Dope dyeing, Yarn dyeing, Piece dyeing, Beck dyeing, Jig dyeing, Padding. **10hrs**

UNIT III: Methods of **printing**-Block, Roller, Screen, tie and dye, Bathik and Modern techniques. **10hrs**

RECOMMENDED TEXT BOOK:

1. Textiles- Fibre to Fabrics- Dr. Bernard, P. Lorbman

WEBSITES & e-LEARNING SOURCES:

1. www.acs.org
2. www.virtlab.com
3. Internet Chemistry resources
4. MATLAB

ETHIRAJ COLLEGE FOR WOMEN (AUTONOMOUS), CHENNAI-600 008

END SEMESTER EXAMINATION

TEMPLATE OF THE QUESTION PAPER

Common to all UG Core, Elective and Allied Theory

Title of the Course -

Max. Marks: 100

Course code-

Time: 3 hrs

SECTION-A

ANSWER ALL THE QUESTIONS

10 x 2=20

QUESTION NUMBERS 1-10

Definition, statement or very short answers with 2 lines, Choosing 2 question from each of the 5 units in the syllabus.

SECTION B

ANSWER ANY 5 QUESTIONS

5 x 8=40

QUESTION NUMBERS FROM 11-18

[8 Questions to be given with a maximum of 2 subdivisions (a,b) choosing not more than 2 questions from each one of the 5 units in the syllabus]

SECTION C

ANSWER ANY 2 QUESTIONS

2 x 20=40

QUESTION NUMBERS FROM 19-22

4 Questions to be given with a maximum of 4 subdivisions (a,b,c,d), with equal weightage to all the 5 units in the syllabus

NOTE: Equal weightage to be given for all the 5 units in the syllabus.

Section-A=10Q x 2 marks	= 20 marks
Section-B=5Q x 8 marks	= 40 marks
Section C=2Q x 20 marks	= 40 marks
Total	= 100 marks

**ETHIRAJ COLLEGE FOR WOMEN (AUTONOMOUS)
CHENNAI-600 008**

END SEMESTER EXAMINATION

Template of the Question Paper Common to All Non Major Electives

Title of the Course: Non Major Elective 1 - Cosmetology

Title of the Course: Non Major Elective 2- Dyeing and Printing of Textiles

Course Code: CH15/1N/COS & CH15/2N/DPT

**Max. Marks: 50
Time: 2 hrs**

ANSWER ANY TEN QUESTIONS OUT OF TWELVE QUESTIONS

10 x 5=50

NOTE: Equal weightage to be given to the theory units I and II in the syllabus.